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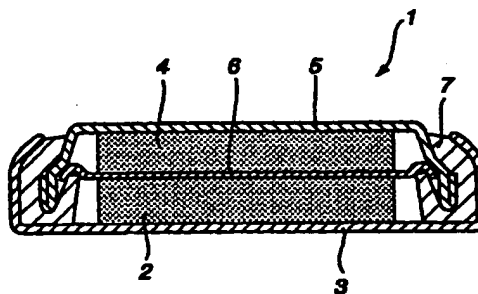
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**(54) METHOD FOR MANUFACTURING ACTIVE MATERIAL OF POSITIVE PLATE AND METHOD FOR MANUFACTURING NONAQUEOUS ELECTROLYTE SECONDARY CELL**

(57) In a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor and a sintering step of sintering and reacting said precursor obtained by said mixing step, in which these substances prove a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals, a reducing agent is added in the above mixing step to said precursor to render it possible to prepare a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.



**FIG.1**

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## Description

## Technical Field

[0001] This invention relates to a method for producing a positive electrode active material that is capable of reversibly doping/undoping lithium, and a method for producing a non-aqueous electrolyte secondary battery employing this positive electrode active material.

## Background Art

[0002] Recently, with the marked progress in a variety of electronic equipment, researches in a rechargeable secondary battery, as a battery that can be used conveniently and economically for prolonged time, are underway. Typical of the known secondary batteries are a lead battery, an alkali storage battery and a lithium secondary battery.

[0003] Of these secondary batteries, a lithium secondary battery has advantages as to high output and high energy density. The lithium secondary battery is made up at least of positive and negative electrodes, containing active materials capable of reversibly introducing and removing lithium ions, and a non-aqueous electrolyte.

[0004] Currently,  $\text{LiCoCO}_2$  is widely exploited as a positive electrode active material of a lithium ion secondary battery having a potential of 4V with respect to the lithium potential. This  $\text{LiCoCO}_2$  is of a high energy density and a high voltage and is an ideal positive electrode material in many respects. However, Co is localized in distribution and represents a rare resource, with the result that, if  $\text{LiCoCO}_2$  is used as a positive electrode active material, the cost is raised, whilst stable supply is difficult.

[0005] Therefore, development of a electrode active material based on Ni, Mn or Fe, that is abundant in supply and inexpensive, is desirable. For example,  $\text{LiNiO}_2$ , based on Ni, has a large theoretical capacity and a high discharging potential. However, in a battery employing  $\text{LiNiO}_2$ , the  $\text{LiNiO}_2$  crystal structure collapses with the charging/discharging cycles, so that the discharging capacity is lowered. On the other hand,  $\text{LiNiO}_2$  suffers a drawback or poor thermal stability.

[0006] As an Mn-based electrode active material, there is proposed  $\text{LiMn}_2\text{O}_4$  having a positive spinel structure and a spatial set Fd3m. This  $\text{LiMn}_2\text{O}_4$  has a high potential of the order of 4V, with respect to the lithium potential, which is equivalent to that of  $\text{LiCoCO}_2$ . Moreover, the  $\text{LiMn}_2\text{O}_4$  is a highly promising material since it is easy to synthesize and has a high battery capacity. However, the battery constructed using  $\text{LiMn}_2\text{O}_4$  is deteriorated in battery capacity since Mn is dissolved in the electrolytic solution with the charging/discharging cycles so that it is insufficient in stability or cycle characteristics.

[0007] On the other hand, such a material having

$\text{LiFeO}_2$  as a basic structure is being researched as an Fe-based electrode active material. Although  $\text{LiFeO}_2$  has a structure similar to that of  $\text{LiCoCO}_2$  or  $\text{LiMn}_2\text{O}_4$ , it is unstable in structure and difficult to synthesize.

[0008] On the other hand, a compound having an ilvitic structure as a positive electrode active material of the lithium secondary battery, such as, for example, a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$ , where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M contains at least one of 3d transition metals (Fe, Mn, Co and Ni), is retained to be a promising material.

[0009] It is proposed in Japanese Laying-Open Patent H-9-171827 to use e.g.,  $\text{LiFePO}_4$ , amongst the compounds represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$ , as a positive electrode of a lithium ion battery. This  $\text{LiFePO}_4$  has a theoretical capacity as large as 170 mAh/g and contains one Li atom, that can be electrochemically doped/undoped in an initial state, per Fe atom, and hence is a promising material as a positive electrode active material for the lithium ion battery.

[0010] Conventionally, this  $\text{LiFePO}_4$  has been synthesized by sintering at a higher temperature of  $800^\circ\text{C}$ , under a reducing environment, using a bivalent iron salt, such as iron phosphate  $\text{Fe}_3(\text{PO}_4)_2$  or iron acetate  $\text{Fe}(\text{CH}_3\text{COO})_2$ , as an Fe source which proves a starting material for synthesis.

[0011] However,  $\text{Fe}^{2+}$  is sensitive to a trace amount of oxygen contained in a synthesizing atmosphere and is readily oxidized to  $\text{Fe}^{3+}$ . The result is that trivalent iron compounds tend to co-exist in the produced  $\text{Fe}_3(\text{PO}_4)_2$  to render it difficult to acquire single-phase  $\text{LiFePO}_4$ .

[0012] It is reported in the above Publication that an actual battery fabricated using the  $\text{Fe}_3(\text{PO}_4)_2$  synthesized by the above-described synthetic method has an actual capacity only as low as approximately 60 mAh/g to 70 mAh/g. Although the actual capacity of the order of 120 mAh/g is subsequently reported in the Journal of the Electrochemical Society, 144,1188 (1997), it cannot be said that a sufficient capacity has been achieved, in consideration that the theoretical capacity is 170 mAh/g.

[0013] If  $\text{LiFePO}_4$  is compared to  $\text{LiMn}_2\text{O}_4$ , the former has a volumetric density of  $3.6 \text{ g/cm}^3$  and an average voltage of 3.4 V; whereas latter has a volumetric density of  $4.2 \text{ g/cm}^3$  and an average voltage of 3.9 V, with the capacity being 120 mAh/g. So,  $\text{LiFePO}_4$  is lower than  $\text{LiMn}_2\text{O}_4$  by approximately 10% in both the voltage and volumetric density. So, for the same capacity of 120 mAh/g,  $\text{LiFePO}_4$  is lower than  $\text{LiMn}_2\text{O}_4$  by not less than 10% and by not less than 20% in weight energy density and in volumetric energy density, respectively. Thus, in order to realize the energy density of  $\text{LiFePO}_4$  which is of the level equivalent to or higher than  $\text{LiMn}_2\text{O}_4$ , the capacity of 140 mAh/g or higher is required. However, this high capacity has not been realized with  $\text{LiFePO}_4$ .

[0014] On the other hand,  $\text{LiFePO}_4$  obtained by the conventional synthesizing method is lower in electrical

conductivity than the positive electrode active material, such as  $\text{LiMn}_2\text{O}_4$ , and hence suffices the problem of low load characteristics and poor molding performance. So, in order to prepare an electrode using  $\text{LiFePO}_4$  as the positive electrode active material, a large quantity of electrically conductive material, such as acetylene black, needs to be added, with the result that a non-aqueous electrolyte secondary battery employing  $\text{LiFePO}_4$  as a positive electrode active material suffers a drawback that the energy density of the electrode is smaller, with the electrode molding performance being also lowered.

#### Disclosure of the Invention

[0015] It is an object of the present invention to provide a method for producing a positive electrode active material in which mixing of an oxidized 3d transition metal into a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  to realize a high capacity and a method for producing a high-capacity non-aqueous electrolyte secondary battery employing this positive electrode active material. It is a particular object of the present invention to provide a method for producing a positive electrode active material in which mixing of an oxidized 3d transition metal into a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  and a method for producing a high-capacity non-aqueous electrolyte secondary battery by employing this positive electrode active material to realize a high capacity in case the compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .

[0016] For accomplishing the above object, the present invention provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor, the substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein a reducing agent is added to the precursor in the mixing step.

[0017] In the method for producing the positive electrode active material according to the present invention, in which the reducing agent is added to the precursor in the mixing step, it is possible to prevent the 3d transition metal  $M$  in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities.

[0018] The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the neg-

ative electrode, in which the positive electrode active material is produced by a method including a mixing step of mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals, to give a precursor, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein a reducing agent is added to the precursor in the mixing step to prepare the positive electrode active material.

[0019] In the method for producing the non-aqueous electrolyte secondary battery according to the present invention, in which the reducing agent is added to the precursor in the mixing step in the preparation of the positive electrode active material, it is possible to prevent the 3d transition metal  $M$  in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities.

[0020] The present invention also provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals, to give a precursor, a de-aerating step of removing air contained in the precursor obtained in the mixing step, a sintering step of sintering and reacting the precursor obtained by the mixing step.

[0021] In the method for producing the positive electrode active material according to the present invention, in which air contained in the precursor in the de-aerating step is removed in the de-aerating step, it is possible to prevent the 3d transition metal  $M$  in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities.

[0022] The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the negative electrode, in which the positive electrode active material is produced by a method including a mixing step of mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals, to give a precursor, a de-aerating step of de-aerating air contained in the precursor obtained in the mixing step and a sintering step of sintering and reacting the precursor obtained in a state free of air by the de-aerating step.

[0023] In the method for producing a non-aqueous

electrolyte secondary battery according to the present invention, in which, in preparing the positive electrode active material, any air contained in the precursor is removed in the deaerating process in the preparation of the positive electrode active material, it is possible to prevent the 3d transition metal M in the starting material for synthesis from being oxidized by the residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities.

[0024] It is another object of the present invention to provide a method for producing a positive electrode active material superior in load characteristics and molding performance, and which is capable of realizing a high capacity, and a method for producing a non-aqueous electrolyte secondary battery which, through the use of this positive electrode active material, has a high capacity approaching to the theoretical capacity of 170 mAh/g.

[0025] For accomplishing this object, the present invention provides a method for producing a positive electrode active material including a mixing step of mixing a plurality of substances to give a precursor, the substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M includes at least one of 3d transition metals, and a sintering step of sintering and reacting the precursor obtained by the mixing step, wherein an electrically conductive agent is added to the starting material for synthesis or to the precursor.

[0026] In the method for producing a positive electrode active material according to the present invention, in which an electrically conductive agent is added to the starting material for synthesis or to the precursor, the produced positive electrode active material exhibits superior load characteristics and electrode molding performance, to realize a high capacity.

[0027] The present invention also provides a method for producing a non-aqueous electrolyte secondary battery having a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing the positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between the positive electrode and the negative electrode, wherein the positive electrode active material is synthesized by a mixing step of mixing a plurality of substances, proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M includes at least one of 3d transition metals, to give a precursor, and a sintering step of sintering and reacting the precursor obtained by the mixing step, and wherein an electrically conductive agent is added to the starting material for synthesis or to the precursor to synthesize the positive electrode active material.

[0028] In the method for producing the non-aqueous electrolyte secondary battery according to the present invention, an electrically conductive agent is

added to the starting material for synthesis or to the precursor. Since the positive electrode active material exhibits satisfactory load characteristics and electrode molding performance, the produced non-aqueous electrolyte secondary battery is of a high capacity.

#### Brief Description of the Drawings

[0029]

Fig.1 is a cross-sectional view showing an illustrative structure of a non-aqueous electrolyte secondary battery embodying the present invention.

Fig.2 is a graph showing a powder X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Example 1.

Fig.3 is a graph showing a powder X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Comparative Example 1.

Fig.4 is a graph showing charging/discharging characteristics of the battery of Example 1.

Fig.5 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 1.

Fig.6 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 1.

Fig.7 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 1.

Fig.8 is a graph showing a powder X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Example 3.

Fig.9 is a graph showing a powder X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Comparative Example 2.

Fig.10 is a graph showing the charging/discharging characteristics of the battery of Example 3.

Fig.11 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 3.

Fig.12 is a graph showing the relation between the sintering temperature in synthesizing  $\text{LiFePO}_4$ .

Fig.13 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 2.

Fig.14 is a graph showing the relation between the

number of cycles and the charging/discharging capacity of the battery of Comparative Example 2.

Fig.15 is a graph showing a powder X-ray diffraction pattern of an  $\text{LiFePO}_4$  carbon complex material synthesized in Example 7.

Fig.16 is a graph showing a powder X-ray diffraction pattern of an  $\text{LiFePO}_4$  carbon complex material synthesized in Example 8.

Fig.17 is a graph showing a powder X-ray diffraction pattern of an  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon complex material synthesized in Example 9.

Fig.18 is a graph showing a powder X-ray diffraction pattern of an  $\text{LiFePO}_4$  carbon complex material synthesized in Comparative Example 7.

Fig.19 is a graph showing the charging/discharging characteristics of the battery of Example 7.

Fig.20 is a graph showing the relation between the number of cycles and the charging/discharging capacity of the battery of Example 7.

Fig.21 is a graph showing load characteristics of a battery of Example 8.

Fig.22 is a graph showing the charging/discharging characteristics of the battery of Example 9.

Fig.23 is a graph showing load characteristics of batteries of Example 7 and Comparative Example 1.

Fig.24 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 4.

Fig.25 is a graph showing powder X-ray diffraction patterns of  $\text{LiFePO}_4$  carbon complex material synthesized in Examples 10 to 12.

Fig.26 is a graph showing powder an X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Comparative Example 5.

Fig.27 is a graph showing the charging/discharging characteristics of the battery of Example 11.

Fig.28 is a graph showing the charging/discharging characteristics of the battery of Example 12.

Fig.29 is a graph showing the charging/discharging characteristics of the battery of Comparative Example 4.

Fig.30 is a graph showing volumetric grain size distribution of the  $\text{LiMnPO}_4$  carbon complex material of Examples 10 to 12 and  $\text{LiMnPO}_4$  of Comparative Example 5.

Fig.31 is a graph showing the volumetric grain size distribution of the  $\text{LiMnPO}_4$  carbon complex material of Example 12 in integrated values of the passed portions.

#### Best mode for Carrying out the Invention

[0030] The present invention is hereinafter explained in detail.

[0031] A non-aqueous electrolyte secondary battery, manufactured in accordance with the present invention, is a so-called lithium secondary battery, and includes, as basic constituent elements, a positive electrode, containing a positive electrode active material, a negative electrode and a non-aqueous electrolyte.

[0032] This positive electrode active material contains a compound having an olivine structure and which is represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$ , where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  contains at least one of 3d transition metals.

[0033] The compounds represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  may be enumerated by, for example,  $\text{Li}_x\text{Fe}_y\text{PO}_4$ ,  $\text{Li}_x\text{Mn}_y\text{PO}_4$ ,  $\text{Li}_x\text{Co}_y\text{PO}_4$ ,  $\text{Li}_x\text{Ni}_y\text{PO}_4$ ,  $\text{Li}_x\text{Cu}_y\text{PO}_4$ ,  $\text{Li}_x(\text{Fe}, \text{Mn})_y\text{PO}_4$ ,  $\text{Li}_x(\text{Fe}, \text{Co})_y\text{PO}_4$ ,  $\text{Li}_x(\text{Fe}, \text{Ni})_y\text{PO}_4$ ,  $\text{Li}_x(\text{Cu}, \text{Mn})_y\text{PO}_4$ ,  $\text{Li}_x(\text{Cu}, \text{Co})_y\text{PO}_4$  and  $\text{Li}_x(\text{Cu}, \text{Ni})_y\text{PO}_4$ , where the proportions of elements in parentheses () are arbitrary.

[0034] The compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  preferably contains particles with the particle size not larger than  $10 \mu\text{m}$ . On the other hand, the specific surface area of the compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  is preferably not less than  $0.5 \text{ m}^2/\text{g}$ .

[0035] In preparing a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  as the aforementioned positive electrode active material, plural substances as starting material for synthesis of the compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  are mixed together to prepare a precursor by a mixing process. The precursor prepared in the mixing process is then sintered and reacted by a sintering process. A reducing agent is added to the precursor in the mixing process.

[0036] The method for preparing  $\text{LiFePO}_4$ , as the compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$ , is now explained.

[0037] In the mixing process, iron oxalate ( $\text{FeC}_2\text{O}_4$ ), as a starting material for synthesis, ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) are mixed together in a pre-set ratio and added to with e.g., iron powders (Fe) as a reducing agent. These substances are mixed sufficiently to give a precursor.

[0038] In the sintering process, this precursor is sin-

tered at a pre-set temperature in an atmosphere of an inert gas, such as nitrogen. This synthesizes  $\text{LiFePO}_4$ .

[0039] If air is left in the precursor in synthesizing  $\text{LiFePO}_4$ ,  $\text{Fe}^{2+}$  in iron oxalate, as a bivalent iron compound, tends to be oxidized with oxygen in air to  $\text{Fe}^{3+}$ . As a result, the trivalent iron compound tends to be mixed as an impurity into the synthesized  $\text{LiFePO}_4$ .

[0040] According to the present invention, iron powders are added as a reducing agent in the mixing process. If  $\text{Fe}^{2+}$  in iron oxalate, as a bivalent iron compound, is oxidized to  $\text{Fe}^{3+}$  by oxygen in air contained in the precursor, the iron powders contained in the precursor reduces this  $\text{Fe}^{3+}$  to  $\text{Fe}^{2+}$ . Since this prohibits the trivalent iron compound from mixing into the synthesized  $\text{LiFePO}_4$ , it is possible to produce single-phase  $\text{LiFePO}_4$ .

[0041] The iron powder, added as a reducing agent, forms a part of the starting material for synthesis, and is synthesized into  $\text{LiFePO}_4$  as an ultimate product. If the reducing agent becomes a part of the starting material for synthesis of  $\text{LiFePO}_4$ , there is no fear of mixing of impurities, so that single-phase  $\text{LiFePO}_4$  may be produced. If the reducing agent is used as a part of  $\text{LiFePO}_4$ , the single-phase  $\text{LiFePO}_4$  may be produced without the risk of mixing of impurities. Moreover, if the reducing agent is used as a part of  $\text{LiFePO}_4$ , there is no necessity of removing the reducing agent left over after the end of the reaction to render it possible to synthesize  $\text{LiFePO}_4$  efficiently.

[0042] If iron powders are used as a part of the reducing agent or of the starting material for  $\text{LiFePO}_4$  synthesis, these iron powders are desirably added to the precursor in an amount ranging from 1 wt% to 30 wt% based on the sum total of Fe in the iron starting material. If the amount of addition of the iron powders is less than this value, oxidation of  $\text{Fe}^{2+}$  cannot be prevented sufficiently. On the other hand, since the iron powders (Fe) are lower in reactivity than  $\text{Fe}^{2+}$  in iron oxalate ( $\text{FeC}_2\text{O}_4$ ), the reaction of synthesis of  $\text{LiFePO}_4$  does not proceed sufficiently if the amount of the iron powders exceeds 30 wt%. By setting the amount of addition of the iron powders in a range from 1 to 30 wt% based on the weight of Fe in the iron starting material, it is possible to prohibit oxidation of  $\text{Fe}^{2+}$  to produce the single-phase  $\text{LiFePO}_4$  without obstructing the synthetic reaction.

[0043] As the reducing agent, oxalic acid, formic acid or hydrogen may be used in place of iron powders, insofar as no solid product is left over after the sintering process.

[0044] In the present technique, in which the reducing agent is added to the precursor as described above to suppress generation of trivalent iron compounds by trace oxygen, and the precursor is sintered in a nitrogen stream, it has become possible to synthesize  $\text{LiFePO}_4$  at a temperature e.g., of  $300^\circ\text{C}$  which is appreciably lower than the conventional temperature of  $800^\circ\text{C}$ . Stated differently,  $\text{LiFePO}_4$  can be synthesized over a

broader temperature range than heretofore to increase the latitude of selection of the precursor sintering temperature, referred to below simply sintering temperature. If the sintering temperature is as high as  $800^\circ\text{C}$ , as in the conventional system, the energy consumption is correspondingly increased, whilst the load imposed on reaction apparatus etc is also increased.

[0045] The present inventors have directed attention to the relation between the precursor sintering temperature in synthesizing  $\text{LiFePO}_4$  and the battery employing  $\text{LiFePO}_4$  as the active material and investigated into an optimum sintering temperature for synthesizing  $\text{LiFePO}_4$  in realization of a high capacity.

[0046] As a result, it has been found that the sintering temperature for sintering the  $\text{LiFePO}_4$  precursor is desirably not lower than  $350^\circ\text{C}$  and not higher than  $790^\circ\text{C}$ . If the sintering temperature is lower than  $350^\circ\text{C}$ , there is a risk that the chemical reaction and crystallization do not proceed sufficiently, such that homogeneous  $\text{LiFePO}_4$  cannot be produced. On the other hand, if the sintering temperature is higher than  $790^\circ\text{C}$ , there is a risk that crystallization proceeds excessively to obstruct lithium diffusion. So, by sintering the precursor at a temperature not lower than  $350^\circ\text{C}$  and not higher than  $790^\circ\text{C}$ , to synthesize  $\text{LiFePO}_4$ , homogeneous  $\text{LiFePO}_4$  can be produced to realize a high capacity exceeding 120 mAh/g which is the capacity of  $\text{LiFePO}_4$  produced by the conventional method.

[0047] It is more desirable that the sintering temperature be not lower than  $450^\circ\text{C}$  and not higher than  $700^\circ\text{C}$ . By sintering the precursor in a temperature range from  $450^\circ\text{C}$  to  $700^\circ\text{C}$ , it is possible to realize the high real capacity approaching to 170 mAh/g which is the theoretical capacity of  $\text{LiFePO}_4$ .

[0048] In the above-described manufacturing method for the positive electrode active material, in which a reducing agent is added to the precursor in synthesizing  $\text{Li}_x\text{M}_y\text{PO}_4$ , it is possible to prevent oxidation of M as a 3d transition metal, so that the single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  can be produced without mixing of impurities. Moreover,  $\text{Li}_x\text{M}_y\text{PO}_4$  can be produced at a sintering temperature lower than in the conventional system. So, with the present manufacturing method for the positive electrode active material,  $\text{Li}_x\text{M}_y\text{PO}_4$  can be produced with which it is possible to realize a high capacity.

[0049] By using  $\text{Li}_x\text{M}_y\text{PO}_4$  synthesized as described above, it is possible to produce a non-aqueous electrolyte secondary battery of high capacity and superior cyclic characteristics and lithium ion doping/undoping performance.

[0050] On the other hand,  $\text{Li}_x\text{M}_y\text{PO}_4$  as a positive electrode active material can be produced as follows: First, plural materials as a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  are mixed to give a precursor by a mixing process. Then, air contained in the precursor obtained by the mixing process is removed by way of a de-aerating process. The precursor freed of air by the de-aerating

ing process is sintered and reacted by a sintering process to produce  $\text{Li}_x\text{M}_y\text{PO}_4$ .

[0051] Another method for synthesizing, e.g.,  $\text{LiFePO}_4$  as a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  is hereinafter explained.

[0052] In the mixing process, iron acetate ( $\text{Fe}(\text{CH}_3\text{COO})_2$ ), hydrogen ammonium phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) are mixed at a pre-set ratio to give the precursor.

[0053] In the de-aerating process, the precursor is homogenized sufficiently and de-aerated to remove air contained in the precursor. As the de-aerating processing, the atmosphere of the precursor is replaced by vacuum to introduce an inert gas. This removes air contained in the precursor. As another example of the de-aerating process, a solvent with a boiling point not higher than  $250^\circ\text{C}$  is made to co-exist with the precursor to vaporize off the solvent into the inert gas. This removes air contained in the precursor. The solvent boiling at  $250^\circ\text{C}$  or lower may be exemplified by, for example, water and ethanol.

[0054] In the sintering process, the precursor from the de-aerating processing is sintered at a pre-set temperature in an atmosphere of an inert gas, such as nitrogen. This synthesizes  $\text{LiFePO}_4$ .

[0055] If air is left in the precursor at the time of synthesis of  $\text{LiFePO}_4$ ,  $\text{Fe}^{2+}$  in iron acetate as a bivalent iron compound may be oxidized by oxygen in air and thereby turned into  $\text{Fe}^{3+}$ . So, the trivalent iron compound may be mixed as an impurity in the synthesized  $\text{LiFePO}_4$ .

[0056] According to the present invention, the air contained in the precursor is removed by the de-aerating process to prevent oxidation of  $\text{Fe}^{2+}$  contained in iron acetate. In this manner, no trivalent iron compound is mixed into synthesized  $\text{LiFePO}_4$  to make it possible to produce single-phase  $\text{LiFePO}_4$ .

[0057] The sintering temperature, which is the temperature of sintering the precursor in synthesizing  $\text{LiFePO}_4$ , is desirably not lower than  $350^\circ\text{C}$  and not higher than  $790^\circ\text{C}$ , as in the case of the range of the sintering temperature.

[0058] In the above-described manufacturing method for the positive electrode active material, the air contained in the precursor is removed in synthesizing  $\text{Li}_x\text{M}_y\text{PO}_4$ , thus preventing oxidation of M as a 3d transition metal. This makes it possible to produce single-phase  $\text{LiFePO}_4$  free of impurities. Moreover,  $\text{Li}_x\text{M}_y\text{PO}_4$  can be synthesized at a lower sintering temperature. Thus, with the present manufacturing method for the positive electrode active material,  $\text{Li}_x\text{M}_y\text{PO}_4$ , realizing a high capacity, may be produced.

[0059] By employing  $\text{Li}_x\text{M}_y\text{PO}_4$ , synthesized as described above, as the positive electrode active material, it is possible to produce a non-aqueous electrolyte secondary battery of high capacity and superior cyclic characteristics and lithium ion doping/undoping performance.

[0060] Moreover, in synthesizing a compound sam-

ple composed of  $\text{Li}_x\text{M}_y\text{PO}_4$  and an electrically conductive material, as a positive electrode active material containing  $\text{Li}_x\text{M}_y\text{PO}_4$ , plural materials as a starting material for synthesis of  $\text{Li}_x\text{M}_y\text{PO}_4$  are mixed to give a precursor by a mixing process. The precursor obtained by the mixing process is then sintered and reacted by way of a sintering process. At this time, an electrically conductive material is added to the starting material for synthesis or to the precursor.

[0061] This electrically conductive material may be exemplified by carbon, silver or an electrically conductive high polymer material. The carbon black may be exemplified by, for example, graphite, acetylene black or the like carbon black material.

[0062] The electrically conductive material is desirably added in a range of 0.5 to 20 parts by weight to 100 parts by weight of  $\text{Li}_x\text{M}_y\text{PO}_4$ . If the amount of the electrically conductive material is less than 0.5 wt%, it is likely that no sufficient effect can be achieved. If conversely the amount of the electrically conductive material exceeds 20 wt%, the proportion of M as the main partner of the oxidation in the positive electrode active material is low such that the energy density of the non-aqueous electrolyte secondary battery is likely to be lowered.

[0063] Therefore, if, in the positive electrode active material, the electrically conductive material is added in a range of 0.5 to 20 parts by weight to 100 parts by weight of  $\text{Li}_x\text{M}_y\text{PO}_4$ , load characteristics and electrode molding performance are improved, such that the non-aqueous electrolyte secondary battery having this compound sample as the positive electrode active material has a high capacity.

[0064] As a manufacturing method for a compound sample as the positive electrode active material, a manufacturing method for synthesizing the  $\text{LiFePO}_4$  carbon compound material composed of  $\text{LiFePO}_4$  and carbon as the  $\text{Li}_x\text{M}_y\text{PO}_4$  and as the electrically conductive material is hereinafter explained.

[0065] In adding carbon to the precursor of the  $\text{LiFePO}_4$ , iron oxalate ( $\text{FeC}_2\text{O}_4$ ), ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) are sufficiently mixed in the mixing process to give a precursor, which precursor then is calcined at a lower temperature in an atmosphere of an inert gas, such as nitrogen. The calcined precursor and the carbon are mixed together and pulverized. In the sintering process, sintering is made at a preset temperature in an atmosphere of an inert gas, such as nitrogen, to produce a  $\text{LiFePO}_4$  carbon compound material.

[0066] If carbon is added to the starting material for synthesis of  $\text{LiFePO}_4$ , carbon is added and mixed at the outset in the mixing process to a starting material for synthesis composed of iron oxalate ( $\text{FeC}_2\text{O}_4$ ), ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) and the resulting mixture is calcined at a lower temperature in an atmosphere of an inert gas, such as nitrogen. In the sintering process, the calcined

mixture is sintered at a pre-set temperature in the inert gas atmosphere, such as nitro gen, to produce an  $\text{LiF PO}_4$  carbon compound material.

[0067] In the present technique, the starting material for synthesis is mixed thoroughly as described above to prepare the precursor. By thoroughly mixing the starting material for synthesis, the respective components of the starting material are mixed evenly to provide a precursor having an increased number of contact points. This precursor is sintered and synthesized in a nitrogen stream. This makes it possible to synthesize the compound sample by sintering the precursor at a temperature of e.g.,  $300^\circ\text{C}$  which is appreciably lower than  $800^\circ\text{C}$  which is the sintering temperature for synthesizing  $\text{LiFePO}_4$  by the conventional method.

[0068] Stated differently, the  $\text{LiFePO}_4$  carbon compound material can be synthesized over a wide temperature range to provide for a wider temperature selection range for synthesis. This sintering temperature is preferably not less than  $350^\circ\text{C}$  and not higher than  $790^\circ\text{C}$  and more preferably not less than  $450^\circ\text{C}$  and not higher than  $700^\circ\text{C}$ .

[0069] If a compound sample composed of a compound of  $\text{Li}_x\text{M}_y\text{PO}_4$ , where M is Mn the redox potential of which is higher than that of Fe, and an electrically conductive material, for example, a compound material composed of  $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$  and carbon, is to be synthesized as the positive electrode active material, iron oxalate ( $\text{FeC}_2\text{O}_4$ ), ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ), and manganese acetate tetrahydrate ( $\text{Mn}(\text{CH}_3\text{COO})_2 \cdot 4\text{H}_2\text{O}$ ) or manganese carbonate ( $\text{MnCO}_3$ ) are mixed at a pre-set ratio. Otherwise, the technique of synthesizing the  $\text{LiF PO}_4$  carbon compound material is used to prepare  $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$ .

[0070] If an  $\text{Li}_x\text{Mn}_y\text{PO}_4$  carbon compound material, composed of  $\text{Li}_x\text{Mn}_y\text{PO}_4$  and carbon where x is such that  $0 < x \leq 2$  and y is such that  $0.8 \leq y \leq 1.2$ , is to be synthesized as the positive electrode active material, the above-described technique of synthesizing the  $\text{LiF PO}_4$  carbon compound material is followed except that manganese carbonate ( $\text{MnCO}_3$ ), ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) are mixed thoroughly at a pre-set ratio for use as a precursor, to prepare the  $\text{Li}_x\text{Mn}_y\text{PO}_4$  precursor.

[0071] The conventional olivine  $\text{Li}_x\text{Mn}_y\text{PO}_4$  is low in electrical conductivity to render oxidation of Mn difficult. So, the present inventors added carbon to the starting material for synthesis of the olivine  $\text{Li}_x\text{Mn}_y\text{PO}_4$  or its precursor to synthesize the  $\text{Li}_x\text{Mn}_y\text{PO}_4$  carbon compound material. Since oxidation of Mn newly occurs in this  $\text{Li}_x\text{Mn}_y\text{PO}_4$  carbon compound material, it is possible to use this material as the positive electrode active material.

[0072] In the above-described manufacturing method for the positive electrode active material, in which the compound sample composed of  $\text{Li}_x\text{Mn}_y\text{PO}_4$

and the electrically conductive material is to be synthesized, it is possible to synthesize a positive electrode active material having load characteristics and electrode molding performance more desirable than those of the positive electrode active material composed solely of  $\text{LiM}_y\text{PO}_4$ . Moreover, with the positive electrode active material, thus prepared, superior load characteristics and electrode molding performance are achieved, even if the amount of the electrically conductive material newly added at the time of preparing the electrodes is small, thus enabling the use of an electrically conductive agent having a large volumetric density, such as graphite.

[0073] So, the non-aqueous electrolyte secondary battery, containing this compound sample as the positive electrode active material, is of high capacity and superior cyclic characteristics, because migration of electrons occurs smoothly in the electrodes. On the other hand, there is no necessity of adding a new electrically conductive agent to the positive electrode mixture of the non-aqueous electrolyte secondary battery, thus increasing the energy density.

[0074] If, in the non-aqueous electrolyte secondary battery employing the positive electrode active material, synthesized as described above, the electrolyte used in a non-aqueous electrolyte secondary battery 1 is liquid, the non-aqueous electrolyte secondary battery 1 includes a positive electrode 2, a positive electrode can 3, accommodating the positive electrode 2, a negative electrode 4, a negative electrode can 5, accommodating the negative electrode 4, a separator 6 mounted between the positive electrode 2 and the negative electrode 4 and an insulating gasket 7. The non-aqueous electrolyte is charged in the positive electrode 2 and in the negative electrode can 5.

[0075] The positive electrode 2 is comprised of a positive electrode collector of, for example, an aluminum foil, and a layer of a positive electrode active material containing the positive electrode active material, synthesized as described above. As a binder contained in the layer of the positive electrode active material, it is possible to use a known resin material routinely used as a binder for the layer of the positive electrode active material for this type of the non-aqueous electrolyte secondary battery.

[0076] The positive electrode can 3, accommodating the positive electrode 2, serves as an external positive electrode for the non-aqueous electrolyte secondary battery 1.

[0077] The negative electrode 4 is formed e.g., by a foil of metal lithium operating as a negative electrode active material. If a material capable of doping/undoping lithium is used as the negative electrode active material, the negative electrode 4 is the layer of the negative electrode active material containing the negative electrode active material and the negative electrode collector which is formed the layer of the negative electrode active material. As the negative electrode collector



lector, a nickel foil, for example, is used. As the negative electrode active material capable of doping/undoping lithium, metal lithium, lithium alloys, an electrically conductive high polymer material doped with lithium, and a laminated compound, such as a carbon material or a metal oxide, may be used. As a binder contained in the layer of the negative electrode active material, any suitable known binders routinely used as a binder for the layer of the negative electrode active material for this sort of the non-aqueous electrolyte secondary battery may be used.

[0078] The negative electrode can 5, accommodating the negative electrode 4, serves as an external positive electrode for the non-aqueous electrolyte secondary battery 1.

[0079] The separator 6, used for separating the positive electrode and the negative electrode from each other, may be formed of any suitable known materials routinely used as a separator for the layer of the negative electrode active material for this sort of the non-aqueous electrolyte secondary battery. For example, a film of a high molecular material, such as polypropylene, is used. From the relation between conductivity of lithium ions and the energy density, it is necessary that the thickness of the separator 6 be as thin as possible. Specifically, the thickness of the separator 6 of, for example, not more than 50  $\mu\text{m}$ , is desirable.

[0080] The insulating gasket 7 is built into and unified with the negative electrode can 5 and serves for preventing leakage of the non-aqueous electrolytic solution charged into the positive electrode can 3 and the negative electrode can 5.

[0081] As the non-aqueous electrolytic solution, a solution obtained on dissolving the electrolyte in a non-protonic non-aqueous solvent is used.

[0082] The non-aqueous solvent may be exemplified by, for example, propylene carbonate, ethylene carbonate, butylene carbonate, vinylene carbonate,  $\gamma$ -butyrolactone, sulforane, 1,2-dimethoxyethane, 1,2-diethoxyethane, 2-methyltetrahydrofuran, 3-methyl 1,3-dioxorane, methyl propionate, methyl lactate, dimethyl carbonate, diethyl carbonate and dipropyl carbonate. Especially, from voltage stability, cyclic carbonates, such as propylene carbonate or vinylene carbonate, or chain carbonates, such as dimethyl carbonate, diethyl carbonate or dipropyl carbonate, are preferably used. As this non-aqueous solvent, only one type non-aqueous solvent or a mixture of two or more non-aqueous solvents may be used.

[0083] As the electrolyte, dissolved in the non-aqueous solvent, lithium salts, such as  $\text{LiPF}_6$ ,  $\text{LiClO}_4$ ,  $\text{LiAsF}_6$ ,  $\text{LiBF}_4$ ,  $\text{LiCF}_3\text{SO}_3$  or  $\text{LiN}(\text{CF}_3\text{SO}_2)_2$ , may be used. Of these lithium salts,  $\text{LiPF}_6$  or  $\text{LiBF}_4$  may preferably be used.

[0084] The non-aqueous electrolyte secondary battery 1 is manufactured, e.g., by the following method:

[0085] For preparing the positive electrode 2, the positive electrode active material and the binder are dis-

persed in a solvent to prepare a slurried positive electrode mixture. The so-prepared positive electrode mixture is evenly coated on a current collector and dried in situ to prepare a layer of the positive electrode active material to complete the positive electrode 2. For preparing the negative electrode 4, the negative electrode active material and the binder are dispersed in a solvent to prepare a slurried negative electrode mixture. The so-prepared negative electrode mixture then is evenly coated on a current collector and dried in situ to prepare a layer of the negative electrode active material to complete the negative electrode 4. The non-aqueous electrolytic solution is prepared by dissolving an electrolyte salt in a non-aqueous solvent.

[0086] The positive electrode 2 and the negative electrode 4 are accommodated in the positive electrode can 3 and in the negative electrode can 5. The separator 6 formed e.g., by a porous polypropylene film is arranged between the positive electrode 2 and the negative electrode 4. The non-aqueous electrolytic solution is charged into the positive electrode can 3 and the negative electrode can 5. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete the non-aqueous electrolyte secondary battery 1.

[0087] In the above-described embodiment, the non-aqueous electrolyte secondary battery 1 employing the non-aqueous electrolytic solution is used, as an example, as the non-aqueous electrolyte secondary battery. The present invention is, however, not limited to this and may be applied to the use as the non-aqueous electrolyte of a solid electrolyte or a gelled solid electrolyte containing a swelling solvent. The present invention may also be applied to a variety of shapes of the non-aqueous electrolyte secondary batteries, such as a cylindrical shape, a square shape, a coin or a button shape, or to a variety of sizes of the non-aqueous electrolyte secondary battery, such as a thin type or large-sized batteries.

[0088] In the above-described embodiment, the manufacturing method for a positive electrode active material by synthesizing  $\text{LiFePO}_4$ , an  $\text{LiFePO}_4$  carbon compound, a compound sample composed of  $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$  and carbon or a  $\text{Li}_x\text{Mn}_y\text{PO}_4$  carbon compound material composed of  $\text{Li}_x\text{Mn}_y\text{PO}_4$  and carbon, has been explained. However, it is only necessary that the positive electrode active material be based on  $\text{Li}_x\text{M}_y\text{PO}_4$  as the basic composition, such that any element may be added to or used as a substituent in  $\text{Li}_x\text{M}_y\text{PO}_4$  as long as the reaction and crystallization proceed at a sintering temperature not lower than 350°C and not higher than 790°C, as described above.  $\text{Li}_x\text{M}_y\text{PO}_4$  may also suffer from defects.

[0089] Also, in the above-described embodying the present invention of the present invention, a solid-phase reaction of mixing and sintering a powdered material as a starting material for synthesis of the compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  in the preparation of the positive electrode active material is explained.

as an example. The present invention, however, is not limited to this and may be applied to the use of various other methods for chemical synthesis than the solid-phase reaction.

[0090] Meanwhile, a starting material of high reactivity is preferably used as the starting material for synthesis of  $\text{Li}_x\text{M}_y\text{PO}_4$ . For example, as the starting material for synthesis of  $\text{LiFePO}_4$ , various starting materials, such as lithium hydroxide, lithium nitrate, lithium acetate, lithium phosphate, iron (II) phosphate or iron (II) oxide may be used in addition to the above-mentioned compounds.

[0091] In the following, specified Examples and Comparative Examples, embodying the present invention, are explained based on the experimental results.

#### (Experiment 1)

[0092] In an experiment 1, in synthesizing  $\text{LiFePO}_4$  as the positive electrode active material, an Example in which iron powders were added as a reducing agent to the precursor and a Comparative Example in which iron powders were not added to the precursor, were synthesized. Plural non-aqueous electrolyte secondary batteries, employing these positive electrode active materials, were prepared, and evaluation was made of difference in characteristics of the respective positive electrode active materials and the non-aqueous electrolyte secondary batteries.

#### Example 1

[0093] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron powders (Fe), iron oxalate dihydrate ( $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 0.4: 1.6: 2: 1 for thirty minutes in a mortar to give a precursor. This precursor was calcined in a nitrogen atmosphere at  $300^\circ\text{C}$  for 12 hours. The so-calcined precursor was sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0094] The so-prepared  $\text{LiFePO}_4$  was used as the positive electrode active material to prepare a battery. 70 wt% of dried  $\text{LiFePO}_4$ , as the positive electrode active material, 25 wt% of acetylene black, as an electrically conductive material, and 5 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into dimethyl formamide as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at  $100^\circ\text{C}$  for one hour to form a layer of the positive electrode active material. The alu-

minum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material.

[0095] A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved  $\text{LiPF}_6$  at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

[0096] The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and the separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped test cell.

#### Example 2

[0097]  $\text{LiFePO}_4$  was prepared in the same way as in Example 1, except using iron acetate in place of iron oxalate as the starting material for synthesis and mixing iron powders (Fe), iron acetate ( $\text{FeCH}_3\text{COO}_2$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) to a molar ratio of 0.2: 1.8: 2: 1. The so-produced  $\text{LiFePO}_4$  was used as the positive electrode active material to prepare a test cell.

#### Comparative Example 1

[0098]  $\text{LiFePO}_4$  was prepared in the same way as in Example 1, except adding no reducing agent and mixing iron oxalate dihydrate ( $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) to a molar ratio of 2: 2: 1. The so-produced  $\text{LiFePO}_4$  was used as the positive electrode active material to prepare a test cell.

[0099] Then, measurement was made of the powder X-ray diffraction pattern of the  $\text{LiFePO}_4$  prepared by the above-described method. The measurement conditions of the powder X-ray diffraction were as follows:

apparatus used: RIGAKU RINT 2500 rotary counter pair negative electrode

goniometer: vertical type standard, radius 185 mm

counter monochromator: used

filter: not used

slit width

divergent slit (DS) = 1°

receiving slit (RS) = 1°

scattering slit (SS) = 0.15 mm

counter device: scintillation counter

measurement method: reflection method, continuous scan

scanning range:  $2\theta = 10^\circ$  to  $80^\circ$

scanning speed: 4°/minute

[0100] The powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Example 1, is shown in Fig.2, from which it is seen that a single-phase  $\text{LiFePO}_4$  has been obtained since the presence of the impurity other than  $\text{LiFePO}_4$  is not confirmed in the product.

[0101] The precursor was sintered at plural temperature points in a range from  $320^\circ\text{C}$  to  $850^\circ\text{C}$  to prepare  $\text{LiFePO}_4$  in a similar manner. It was found that there was no impurity in  $\text{LiFePO}_4$  synthesized in the range from  $320^\circ\text{C}$  to  $850^\circ\text{C}$  so that a single-phase  $\text{LiFePO}_4$  has been obtained.

[0102] On the other hand, the presence of the impurity other than  $\text{LiFePO}_4$  was not noticed in the product synthesized in Example 2 from the powder X-ray diffraction pattern of  $\text{LiFePO}_4$  synthesized in Example 2, thus testifying to the generation of the single-phase  $\text{LiFePO}_4$ .

[0103] The powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Comparative Example 2, is shown in Fig.3, from which it is seen that impurities other than  $\text{LiFePO}_4$  are present in the product such that no single-phase  $\text{LiFePO}_4$  has been yielded. Meanwhile, in Fig.3, a peak attributable to the impurity is marked with an asterisk \*. This impurity is presumably a trivalent iron compound, such as  $\text{Fe}_2\text{O}_3$ .

[0104] The test cells, prepared as described above, were subjected to the charging/discharging test, in which each test cell was charged by constant current charging and, when the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below  $0.01 \text{ mA/cm}^2$ . The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V. Meanwhile, charging/discharging was carried out at ambient temperature ( $23^\circ\text{C}$ ), with the current density at this time being  $0.12 \text{ mA/cm}^2$ .

[0105] The charging/discharging characteristics of the battery of Example 1 are shown in Fig.4, from which it is seen that the battery of Example 1 showed a flat potential in the vicinity of 3.4 V, thus indicating that the high reversible charging/discharging capacity of 163

$\text{mAh/g}$ , which is close to the theoretical capacity of 170  $\text{mAh/g}$ , is being produced.

[0106] The relation between the number of cycles and the charging/discharging capacity of the battery of Example 1 is shown in Fig.5, from which it is seen that the battery of the Example 1 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

[0107] The charging/discharging characteristics of the battery of Comparative Example 1 are shown in Fig.6. The relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 1 is shown in Fig.7. From Figs.6 and 7, it may be seen that the battery of Comparative

Example 1 is satisfactory in cyclic characteristics, however, the battery of Comparative Example 1 suffers from significant capacity loss at an initial stage, with the charging/discharging efficiency being low.

[0108] Thus, it may be seen that, by adding a reducing agent to the precursor in synthesizing  $\text{LiFePO}_4$ ,  $\text{Fe}^{2+}$  in the precursor may be prevented from being oxidized by the residual oxygen to become  $\text{Fe}^{3+}$  to give a single-phase  $\text{LiFePO}_4$ . The battery employing this single-phase  $\text{LiFePO}_4$  as the positive electrode active material is superior in charging/discharging characteristics and in cyclic characteristics.

#### (Experiment 2)

[0109] In this experiment 2, in synthesizing  $\text{LiFePO}_4$  as the positive electrode active material, an Example in which a precursor was de-aerated, and a Comparative Example in which iron powders were not added, were synthesized, and plural non-aqueous electrolyte secondary batteries employing these positive electrode active materials were prepared. Of the respective positive electrode active materials and the non-aqueous electrolyte secondary batteries, the difference in characteristics was evaluated.

#### Example 3

[0110] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ( $\text{FeCH}_3\text{COO})_2$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor.

[0111] This precursor then was charged into an electrical oven and, after the atmosphere was replaced by vacuum, a nitrogen gas was introduced into the oven. This precursor was calcined in a nitrogen atmosphere at  $300^\circ\text{C}$  for 12 hours, after which the precursor was sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0112] The so-prepared  $\text{LiFePO}_4$  was used as a

positive electrode active material to produce a test cell in the same way as in Example 1.

#### Example 4

[0113] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ( $\text{FeCH}_3\text{COO})_2$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2mm in diameter, with water as a solvent.

[0114] This mixture then was dried to a clayey state, after which the mixture was sintered in a nitrogen stream in an electrical oven at  $120^\circ\text{C}$  to remove residual water to produce a precursor with high homogeneity. This precursor then was calcined at  $300^\circ\text{C}$  for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0115] Using the so-prepared  $\text{LiFePO}_4$  as the positive electrode active material, a test cell was prepared as in Example 1.

#### Example 5

[0116] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ( $\text{FeCH}_3\text{COO})_2$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2 mm in diameter, with ethanol as a solvent.

[0117] This mixture then was dried to a clayey state, after which the mixture was sintered in a nitrogen stream in an electrical oven at  $120^\circ\text{C}$  to remove residual ethanol to produce a precursor with extremely high homogeneity. This precursor then was calcined at  $300^\circ\text{C}$  for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0118] Using the so-prepared  $\text{LiFePO}_4$  as the positive electrode active material, a test cell was prepared as in Example 1.

#### Example 6

[0119] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) as a starting material of a coarser crystallite size was sufficiently pulverized at the outset. Then, iron acetate ( $\text{FeCH}_3\text{COO})_2$ , ammonium dihydrogen

phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting mixture was further mixed for 48 hours, using zirconia balls 2 mm in diameter, with acetone as a solvent.

[0120] This mixture then was dried to a clayey state, after which the mixture was sintered in a nitrogen stream in an electrical oven at  $120^\circ\text{C}$  to remove residual acetone to produce a precursor with high homogeneity. This precursor then was calcined at  $300^\circ\text{C}$  for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0121] Using the so-prepared  $\text{LiFePO}_4$  as the positive electrode active material, a test cell was prepared as in Example 1.

#### Comparative Example 2

[0122] First,  $\text{LiFePO}_4$  was synthesized. For synthesizing  $\text{LiFePO}_4$ , iron acetate ( $\text{FeCH}_3\text{COO})_2$ , ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 2: 2: 1 for thirty minutes in a mortar to give a precursor. The resulting precursor was calcined at  $300^\circ\text{C}$  for 12 hours in a nitrogen atmosphere and further sintered in the nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ . Using the so-prepared  $\text{LiFePO}_4$  as the positive electrode active material, a test cell was prepared as in Example 1.

[0123] Of the  $\text{LiFePO}_4$ , synthesized as described above, measurement was made of the powder X-ray diffraction pattern under the above-mentioned measurement conditions.

[0124] Fig.8 shows the powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Example 3. It is seen from Fig.8 that no impurity other than  $\text{LiFePO}_4$  was confirmed to exist in the product so that the single-phase  $\text{LiFePO}_4$  has been produced.

[0125] Also, from the powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Examples 4 to 6, no impurity other than  $\text{LiFePO}_4$  was confirmed to exist in the product, so that, in these Examples, the single-phase  $\text{LiFePO}_4$  has been produced.

[0126] The powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Comparative Example 2, is shown in Fig.9, from which it is seen that impurities other than  $\text{LiFePO}_4$  are present in the product such that no single-phase  $\text{LiFePO}_4$  has been yielded. Meanwhile, in Fig.3, the peak attributable to the impurity is marked with an asterisk \*. This impurity is presumably a trivalent iron compound, such as  $\text{Fe}_2\text{O}_3$ .

[0127] The test cells, prepared as described above, were subjected to the charging/discharging test, using the same method as stated in Experiment 1.

[0128] Of the battery of Example 3, the charging/discharging characteristics for the first to third cycles are shown in Fig.10, from which it is seen that the bat-

tery of Example 3 showed a flat potential in the vicinity of 3.4 V, thus indicating that the battery has a high reversible charging/discharging capacity of 163 mAh/g, which is close to the theoretical capacity of 170 mAh/g.

[0129] The relation between the number of cycles and the charging/discharging capacity of the battery of Example 3 is shown in Fig. 11, from which it is seen that the battery of the Example 3 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

[0130] Similarly to the battery of Example 3, the batteries of Examples 4 to 6 were confirmed to exhibit a high charging/discharging capacity and superior cyclic characteristics.

[0131] Test cells were prepared in the same way as in Example 6 except changing the sintering temperature of the precursor to check into the charging/discharging capacity of the cells. The relation between the sintering temperature of the precursor and the charging/discharging capacity of the cells is shown in Fig. 12, from which it is seen that, by synthesizing  $\text{LiFePO}_4$  at a temperature not lower than 350°C and not higher than 790°C, a high capacity exceeding 120 mAh/g of a conventional non-aqueous electrolyte secondary battery can be achieved.

[0132] The charging/discharging characteristics of the battery of Comparative Example 2 are shown in Fig. 13. The relation between the number of cycles and the charging/discharging capacity of the battery of Comparative Example 2 is shown in Fig. 14. It is seen from Figs. 13 and 14 that the battery of Comparative Example 2 is satisfactory in the cyclic characteristics, however, it is high in capacity loss in an initial state and low in the charging/discharging efficiency.

[0133] Thus, it is seen that, by de-aerating the precursor in synthesizing  $\text{LiFePO}_4$ , it is possible to prevent  $\text{Fe}^{2+}$  in the precursor from being oxidized with residual oxygen to  $\text{Fe}^{3+}$  to yield single-phase  $\text{LiFePO}_4$ . It is also seen that the battery employing this single-phase  $\text{LiFePO}_4$  as the positive electrode active material is superior in charging/discharging characteristics and cyclic characteristics.

#### (Experiment 3-1)

[0134] In an experiment 3-1, a variety of positive electrode active materials, containing  $\text{LiFePO}_4$  or  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$ , were synthesized, and non-aqueous electrolyte secondary batteries, employing these positive electrode active materials, were prepared, and evaluation was made of the difference in characteristics of these non-aqueous electrolyte secondary batteries.

#### Example 7

[0135] An  $\text{LiFePO}_4$  carbon compound material composed of  $\text{LiFePO}_4$  as a positive electrode active material and carbon as an electrically conductive mate-

rial was synthesized and, using this compound material, non-aqueous electrolyte secondary batteries were prepared.

[0136] First, iron oxalate dihydride ( $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at 300°C for two hours to yield a precursor. This precursor and graphite were mixed to a weight ratio of 90:10 and pulverized sufficiently by a ball mill for mixing. The precursor then was sintered in a nitrogen atmosphere at 550°C for 24 hours to synthesize an  $\text{LiFePO}_4$  carbon compound material.

[0137] A battery was prepared, using the  $\text{LiFePO}_4$  carbon compound material, obtained as described above, as a positive electrode active material first, 97 wt% of the dried  $\text{LiFePO}_4$  carbon compound material and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at 100°C for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved  $\text{LiPF}_6$  at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

[0138] The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The electrode cans 3, 5 are caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery as a test cell.

#### Example 8

[0139] An  $\text{LiFePO}_4$  precursor was prepared in the same way as in Example 7. A  $\text{LiFePO}_4$  carbon compound material was prepared in the same way as in Example 7 except mixing this precursor, graphite and acetylene black together to a weight ratio of 90: 5: 5 and sufficiently pulverizing the mixture by a ball mill for mix-

ing.

[0140] Using the  $\text{LiFePO}_4$  carbon compound material, obtained as described above, as a positive electrode active material, a test cell was prepared in the same way as in Example 7.

#### Example 9

[0141] A  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material was prepared as the positive electrode active material and a test cell was prepared using the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material as the positive electrode active material.

[0142] First, iron oxalate dihydride ( $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ), manganese acetate tetrahydride ( $\text{Mn}(\text{CH}_3\text{COO})_2$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed together to a molar ratio of 1: 1: 2: 1. The resulting mixture was sufficiently pulverized by a ball mill for mixing. The resulting mixture was calcined in a nitrogen atmosphere at  $300^\circ\text{C}$  for two hours to give a precursor. This precursor was mixed with graphite and acetylene black to give a weight ratio of 90 : 6: 4 and sufficiently pulverized by a ball mill for mixing. The so-calcined precursor was sintered in the nitrogen atmosphere at  $450^\circ\text{C}$  for 24 hours to synthesize the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material.

[0143] Using the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material, obtained as described above, as a positive electrode active material, a test cell was prepared in the same way as in Example 7.

#### Comparative Example 3

[0144] As Comparative Example 3,  $\text{LiFePO}_4$  was synthesized as a positive electrode active material and a test cell was fabricated using  $\text{LiFePO}_4$  as the positive electrode active material.

[0145] First, iron oxalate dihydride ( $\text{FeC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed to a molar ratio of 2: 2: 1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at  $300^\circ\text{C}$  for two hours to yield a precursor. This precursor was sintered in a nitrogen atmosphere at  $550^\circ\text{C}$  for 24 hours to synthesize  $\text{LiFePO}_4$ .

[0146] A battery was prepared, using the  $\text{LiFePO}_4$ , obtained as described above, as a positive electrode active material. First, 87 wt% of the dried  $\text{LiFePO}_4$ , 10 wt% of graphite, as an electrically conductive agent, and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at

$100^\circ\text{C}$  for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved  $\text{LiPF}_6$  at a concentration of 1 mol/l to prepare a non-aqueous electrolytic solution.

[0147] The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolytic solution was charged into the positive electrode can and the negative electrode can. The positive and negative electrode cans were caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery.

[0148] Of the compound samples, synthesized in the Examples 7 to 9, and  $\text{LiFePO}_4$ , prepared in the Comparative Example 3, the powder X-ray diffraction pattern was measured under the same measurement conditions shown in Experiment 1 above. The measured results of the powder X-ray diffraction are shown in Figs.15 to 18.

[0149] The powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Example 7, is shown in Fig.15, from which it is seen that no impurities other than  $\text{LiFePO}_4$  are confirmed in the product such that the product yielded is single-phase  $\text{LiFePO}_4$ . Meanwhile, in Fig.3, the peak attributable to the impurity is marked with an asterisk \*.

[0150] The precursor was sintered at plural temperature points in a range from  $320^\circ\text{C}$  to  $850^\circ\text{C}$  to prepare the  $\text{LiFePO}_4$  carbon compound material in the same manner as in Example 7. It was found that there was no impurity present in the  $\text{LiFePO}_4$  carbon compound material thus testifying to the yielding of the single-phase  $\text{LiFePO}_4$ .

[0151] The powder X-ray diffraction pattern of the  $\text{LiFePO}_4$  carbon compound material synthesized in Example 8 is shown in Fig. 16, from which it is seen that the presence of impurities other than  $\text{LiFePO}_4$  is not confirmed in the product, except the diffraction peak of graphite at approximately  $26^\circ$ , thus testifying to the yielding of the single-phase  $\text{LiFePO}_4$ . In Fig.16, the peak attributable to graphite is indicated by an asterisk \*.

[0152] The powder X-ray diffraction pattern of the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material, synthesized in Example 9, is shown in Fig.9. The powder X-ray diffraction patterns of the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound materials, synthesized with the sintering temperature

atures of 500°C, 600°C and 700°C, are also shown in Fig.17. It may be seen from Fig.17 that, in the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound materials, synthesized at several temperature points in the range from 450°C to 850°C, the presence of impurities other than  $\text{LiFePO}_4$  is not confirmed, except the diffraction peak of graphite at approximately  $26^\circ$ , thus testifying to the yielding of the single-phase  $\text{LiFePO}_4$ . In Fig. 16, the peak attributable to graphite is indicated by an asterisk.

[0153] The powder X-ray diffraction pattern of  $\text{LiFePO}_4$ , synthesized in Comparative Example 3, is shown in Fig.18, from which it is seen that no impurities other than  $\text{LiFePO}_4$  are confirmed to exist in the product such that the product yielded is single-phase  $\text{LiFePO}_4$ .

[0154] The following charging/discharging test was conducted on the test cells of the Examples 7 to 9 and the Comparative Example 3, prepared as described above, to evaluate battery characteristics.

[0155] As this charging/discharging test, each test cell was charged by constant current charging and, when the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below  $0.05 \text{ mA/cm}^2$ . The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V.

[0156] Both charging and discharging were carried out at ambient temperature ( $25^\circ\text{C}$ ) and, in evaluating charging/discharging characteristics, the current density was set to  $0.12 \text{ mA/cm}^2$ . In evaluating load characteristics, the charging/discharging test was conducted under various different current densities. The results of the above charging/discharging test are shown in Figs.19 to 23.

[0157] Of the battery of Example 7, the charging/discharging characteristics for the first to third cycles are shown in Fig.19, from which it is seen that the battery of Example 7 showed a flat potential in the vicinity of 3.4 V, thus indicating that the battery has a high reversible charging/discharging capacity of  $155 \text{ mAh/g}$ , which is close to the theoretical capacity of  $170 \text{ mAh/g}$ .

[0158] The relation between the number of cycles and the charging/discharging capacity of the battery of Example 7 is shown in Fig.20, from which it is seen that the battery of the Example 7 undergoes cyclic deterioration as low as 0.1%/cycle, thus demonstrating stable battery characteristics.

[0159] The load characteristics of the battery of Example 8 are shown in Fig.21, from which it is seen that, if charged/discharged at 0.2 mA, the battery of Example 8 gives a high capacity of  $163 \text{ mAh/g}$ , and that the reversible charging/discharging capacity for the charging/discharging current density of  $2 \text{ mA/cm}^2$  is kept at 93% of that for the charging/discharging current density of  $0.2 \text{ mA/cm}^2$ .

[0160] The charging/discharging characteristics of the battery of Example 9 are shown in Fig.22, from which it is seen that the battery of Example 9 has the reversible charging/discharging capacity of  $146 \text{ mAh/g}$  close to the theoretical capacity of  $170 \text{ mAh/g}$ , with the average discharging potential being 3.58 V.

[0161] The load characteristics of the batteries of Example 7 and Comparative Example 3 are shown in Fig.23, from which it is seen that the reversible charging/discharging capacity for the charging/discharging current density of  $1 \text{ mA/cm}^2$  is kept at 90% of that for the charging/discharging current density of  $0.2 \text{ mA/cm}^2$ . In the battery of the Comparative Example 3, the same amount of graphite as that used as the electrically conductive agent in Example 7 is added as the electrically conductive agent in the positive electrode mixture. However, the battery of the Comparative Example 3 is low in battery capacity and inferior in load characteristics as compared to the battery of Example 7.

[0162] The charging/discharging curve of the battery employing  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  solely as the positive electrode mixture according to the teaching of the publication J. Electrochem. Soc. 144,1188 (1997), is shown in Fig.11. Meanwhile, the battery employing  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  solely as the positive electrode mixture is the battery of Comparative Example 4. It is seen from Fig.24 that the battery of the Comparative Example 4 has the battery capacity as low as  $75 \text{ mAh/g}$ .

[0163] The above results indicate that the batteries of Examples 7 and 8 are non-aqueous electrolyte secondary batteries containing  $\text{LiFePO}_4$  and carbon as the positive electrode active material and as the electrically conductive material, respectively, and hence are superior in load characteristics and of high capacity as compared to the battery of the Comparative Example 3 employing solely the compound of the general formula  $\text{LiFePO}_4$  as the positive electrode mixture. On the other hand, the battery of Example 9 is the non-aqueous electrolyte secondary battery containing  $\text{LiFePO}_4$  and carbon as the positive electrode active material and as the electrically conductive material, respectively, and hence is superior in load characteristics and of high capacity as compared to the battery of the Comparative Example 4 containing solely the compound represented by the general formula  $\text{LiFePO}_4$  and carbon as the compound sample as the positive electrode active material and as the electrically conductive agent, respectively.

[0164] It is also seen from Figs.23 and 21 that, on comparing Examples 7 and 8, the battery of Example 8 is superior in load characteristics to the battery of Fig.7 in which graphite alone is added to the precursor. From this it is seen that, with the non-aqueous electrolyte secondary battery, load characteristics may be improved by employing graphite and acetylene black in combination as carbon in synthesizing the positive electrode active material.

[0165] Comparison of the Examples 7 and 8 in Figs.19 and 22 also reveals that the battery of Example

9 has a higher value of the discharging potential. From this it is seen that, by employing the  $\text{LiFe}_{0.5}\text{Mn}_{0.5}\text{PO}_4$  carbon compound material as the positive electrode active material, it is possible to realize a non-aqueous electrolyte secondary battery of higher capacity and higher voltage value.

#### (Experiment 3-2)

[0166] In the experiment 3-2, a variety of positive electrode active materials containing  $\text{LiMnPO}_4$  were synthesized and non-aqueous electrolyte secondary batteries containing these positive electrode active materials were prepared, and evaluation was made of difference in characteristics of the respective non-aqueous electrolyte secondary batteries.

#### Example 10

[0167] As a positive electrode active material,  $\text{LiFePO}_4$  carbon compound material was synthesized and a test cell containing the  $\text{LiFePO}_4$  carbon compound material as the positive electrode active material was prepared.

[0168] First, manganese carbonate ( $\text{MnCO}_3$ ), ammonium dihydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed to a molar ratio of 2:2:1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at  $280^\circ\text{C}$  for three hours to yield a precursor. This precursor and acetylene black were mixed to a weight ratio of 95:5 and pulverized sufficiently by a ball mill for mixing. The precursor then was sintered in a nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize a  $\text{LiFePO}_4$  carbon compound material.

[0169] A battery was prepared, using the  $\text{LiFePO}_4$ , obtained as described above, as a positive electrode active material. First, 97 wt% of the dried  $\text{LiFePO}_4$  carbon compound material, 10 wt% of graphite, and 3 wt% of polyvinylidene fluoride, as a binder, were evenly mixed into N-methyl pyrrolidone as a solvent to prepare a paste-like positive electrode mixture. Meanwhile, #1300 manufactured by Aldrich Inc. was used as the polyvinylidene fluoride. This positive electrode mixture was applied to an aluminum mesh, as a current collector, and dried in situ in a dry argon atmosphere at  $100^\circ\text{C}$  for one hour to form a layer of the positive electrode active material. The aluminum mesh, on which the layer of the positive electrode active material was formed, was punched to a disc 15.5 mm in diameter to form a pellet-like positive electrode. Meanwhile, this positive electrode carries 60 mg of the active material. A metal lithium foil was punched to substantially the same shape as the positive electrode and used as a negative electrode. In a mixed solvent of equal parts in volume of propylene carbonate and dimethyl carbonate was dissolved  $\text{LiPF}_6$  at a concentration of 1 mol/l to prepare a non-aqueous electrolyte solution.

[0170] The positive electrode, prepared as described above, was accommodated in the positive electrode can, whilst the negative electrode was accommodated in the negative electrode can and a separator was arranged between the positive electrode and the negative electrode. The non-aqueous electrolyte solution was charged into the positive electrode can and the negative electrode can. The positive and negative electrode cans were caulked fixedly through the insulating gasket 7 to complete a 2025 type coin-shaped non-aqueous electrolyte secondary battery.

#### Example 11

[0171] An  $\text{LiFePO}_4$  carbon compound material was prepared in the same way as in Example 10 except mixing the precursor and acetylene black in a weight ratio of 90:10. Using this  $\text{LiFePO}_4$  carbon compound material as the positive electrode mixture, a test cell was prepared in the same way as in Example 10.

#### Example 12

[0172] The  $\text{LiFePO}_4$  carbon compound material was synthesized in the same way as in Example 10 to synthesize  $\text{LiFePO}_4$  carbon compound material. Using this  $\text{LiFePO}_4$  carbon compound material as the positive electrode mixture, a test cell was prepared in the same way as in Example 10.

#### Comparative Example 5

[0173] As a positive electrode active material,  $\text{LiMnPO}_4$  was synthesized and a test cell containing the  $\text{LiFePO}_4$  carbon compound material as the positive electrode active material was prepared.

[0174] First, manganese carbonate ( $\text{MnCO}_3$ ), ammonium hydrogen phosphate ( $\text{NH}_4\text{H}_2\text{PO}_4$ ) and lithium carbonate ( $\text{Li}_2\text{CO}_3$ ) were mixed to a molar ratio of 2:2:1 and thoroughly pulverized and mixed by a ball mill. The resulting mixture was calcined in a nitrogen atmosphere at  $300^\circ\text{C}$  for three hours to yield a precursor. This precursor then was sintered in a nitrogen atmosphere at  $600^\circ\text{C}$  for 24 hours to synthesize  $\text{LiMnPO}_4$ .

[0175] For preparing a battery employing  $\text{LiMnPO}_4$ , obtained as described above, as the positive electrode active material, a test cell was prepared in the same way as in Example 10, except evenly mixing 85 wt% of dried  $\text{LiMnPO}_4$  as a positive electrode mixture, 10 wt% of graphite as an electrically conductive agent and 10 wt% of polyvinylidene fluoride as a binder, in N-methyl pyrrolidone, as a solvent, to prepare a paste-like positive electrode mixture.

[0176] Of the compound samples of Examples 10 to 12, synthesized as described above, and  $\text{LiMnPO}_4$  of Comparative Example 5, a powder X-ray diffraction pattern was measured in accordance with the method and conditions described above.



[0177] The powder X-ray diffraction pattern of the  $\text{LiMnPO}_4$  carbon compound material, synthesized in accordance with Examples 10 to 12, is shown in Fig.25, from which it is seen that no impurities other than  $\text{LiMnPO}_4$  is confirmed to exist in the product, such that the product yielded is single-phase  $\text{LiFePO}_4$ .

[0178] The powder X-ray diffraction pattern of the  $\text{LiMnPO}_4$  carbon compound material, synthesized in accordance with Comparative Example 5, is shown in Fig.26, from which it is seen that no impurities other than  $\text{LiMnPO}_4$  is confirmed to exist in the product such that the product yielded is single-phase  $\text{LiFePO}_4$ .

[0179] The following charging/discharging test was conducted on the test cells of the Example 11 and the Comparative Example 5, prepared as described above, to evaluate battery characteristics. As this charging/discharging test, each test cell was charged by constant current charging with the current of  $0.5 \text{ mA/cm}^2$ . When the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.5 V was kept. The charging was stopped when the current fell below  $0.05 \text{ mA/cm}^2$ . The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V.

[0180] The following charging/discharging test was conducted on the test cell of the Example 12, prepared as described above, to evaluate battery characteristics. As this charging/discharging test, the test cell was charged by constant current charging with the current of  $0.5 \text{ mA/cm}^2$ . When the battery voltage reached 4.5V, the charging system was switched from the constant current charging to constant voltage charging, and charging was carried out as the voltage of 4.8 V was kept. The charging was stopped when the current fell below  $0.05 \text{ mA/cm}^2$ . The discharging then was carried out and stopped at a time point when the battery voltage was lowered to 2.0 V. Meanwhile, the battery charging and discharging were carried out at an ambient temperature ( $23^\circ\text{C}$ ). The results obtained by this charging/discharging test are shown in Figs.27 to 29.

[0181] Fig.27 shows a graph representing charging/discharging characteristics for the battery of Example 11. It is seen from Fig.27 that the battery of Fig.11 has a flat potential in the vicinity of 4V and generates the reversible charging/discharging capacity of 113  $\text{mAh/g}$ . On the other hand, the charging/discharging characteristics of the battery of Example 12 are shown in a graph of Fig.28, from which it is seen that the discharging voltage is high, with the discharging capacity being as large as 120  $\text{mAh/g}$ . Conversely, the battery of Comparative Example 5 is free of a flat discharging area, with there occurring no Mn oxidation.

[0182] It may be seen from the above results that the batteries of Examples 11 and 12 are non-aqueous electrolyte secondary batteries obtained by using a synthesized compound sample composed of  $\text{LiMnPO}_4$  as a positive electrode mixture and of acetylene black as an

electrically conductive agent, such that these batteries are superior in load characteristics and of high capacity as compared to the battery of Comparative Example 5 employing  $\text{LiFePO}_4$  as the positive electrode active material.

[0183] From measuring the volume grain size distribution of the  $\text{LiMnPO}_4$  carbon compound material of Examples 10 to 12 and  $\text{LiMnPO}_4$  of Comparative Example 5, the scattering of the laser light was measured using a volume grain size distribution measurement device, manufactured by HORIBA SEISAKUSHO CO. LTD. under the trade name of Micro-Lack grain size analyzer LA-920. The measured results of the volume grain size distribution are shown in Fig.30, from which it is seen that the  $\text{LiMnPO}_4$  carbon compound material and  $\text{LiMnPO}_4$  all contain particles not larger than  $10 \mu\text{m}$ .

[0184] By representing the volume grain size distribution of the  $\text{LiMnPO}_4$  carbon compound material of Example 12 by integrated amount of passed grains, it has been seen that the grain size in its entirety is not more than  $6 \mu\text{m}$ . That is, by adding carbon to the synthetic material of  $\text{LiMnPO}_4$  or to the precursor, the grain growth of the precursor during the sintering process can be suppressed to realize more homogeneous finer compound sample.

#### Industrial Applicability

[0185] As may be seen from the foregoing description, the manufacturing method of the positive electrode mixture according to the present invention adds a reducing agent to the precursor during the mixing process. Since this prevents the 3d transition metal M in the synthetic starting material from being oxidized by residual oxygen, thus yielding a single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities, it is possible to prepare a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.

[0186] Moreover, in the manufacturing method for the non-aqueous electrolyte secondary battery according to the present invention, a reducing agent is added to the precursor during the mixing process. This prevents the 3d transition metal M in the synthetic starting material from being oxidized by residual oxygen to yield a single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities. This  $\text{Li}_x\text{M}_y\text{PO}_4$  renders it possible to produce a non-aqueous electrolyte secondary battery of high capacity because lithium is diffused sufficiently in the grains.

[0187] Moreover, with the manufacturing method of the positive electrode active material according to the present invention, air contained in the precursor is removed in the de-aerating process. This prevents the 3d transition metal M in the starting material for synthesis from being oxidized by residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities, thus enabling the manufacture of a positive electrode active material capable of reversibly and satisfactorily doping/undoping lithium.

[0188] Moreover, with the manufacturing method of the positive electrode active material according to the present invention, air contained in the precursor is removed in the de-aerating process in preparing the positive electrode active material. This prevents the 3d transition metal M in the starting material for synthesis from being oxidized by residual oxygen to yield single-phase  $\text{Li}_x\text{M}_y\text{PO}_4$  free of impurities. With this  $\text{Li}_x\text{M}_y\text{PO}_4$ , since lithium diffusion in the particle occurs sufficiently, it becomes possible to produce a high-capacity non-aqueous electrolyte secondary battery.

[0189] In addition, in the manufacturing method for the positive electrode active material according to the present invention, an electrically conductive agent is added to the starting material for synthesis or to the precursor of the positive electrode active material. The positive electrode active material produced exhibits superior load characteristics and electrode molding performance thus realizing a high capacity.

[0190] Moreover, in the manufacturing method for the positive electrode active material according to the present invention, an electrically conductive agent is added to the starting material for synthesis or to the precursor in synthesizing the positive electrode active material. Since the positive electrode active material produced thus exhibits superior load characteristics and electrode molding performance, the non-aqueous electrolyte secondary battery produced exhibits high capacity as the non-aqueous electrolyte secondary battery.

#### Claims

1. A method for producing a positive electrode active material comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M includes at least one of 3d transition metals; and

a sintering step of sintering and reacting said precursor obtained by said mixing step; wherein a reducing agent is added to said precursor in said mixing step.

2. The method for producing a positive electrode active material according to claim 1 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .
3. The method for producing a positive electrode active material according to claim 2 wherein said reducing agent is Fe which is also a part of said starting material for synthesis.
4. A method for producing a non-aqueous electrolyte

secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode, said positive electrode active material being produced by a method comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M includes at least one of 3d transition metals; and

a sintering step of sintering and reacting said precursor obtained by said mixing step; wherein a reducing agent is added to said precursor in said mixing step to prepare said positive electrode active material.

5. The method for producing a non-aqueous electrolyte secondary battery according to claim 4 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .
6. The method for producing a non-aqueous electrolyte secondary battery according to claim 5 wherein said reducing agent is Fe which is also a part of said starting material for synthesis.
7. A method for producing a positive electrode active material comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$ , y is such that  $0.8 \leq y \leq 1.2$  and M includes at least one of 3d transition metals;

a de-aerating step of removing air contained in said precursor obtained in said mixing step;

a sintering step of sintering and reacting said precursor obtained by said mixing step.

8. The method for producing a positive electrode active material according to claim 7 wherein an inert gas is introduced after substituting vacuum for an atmosphere in said precursor and subsequently an inert gas is introduced to remove air contained in said precursor.
9. The method for producing a positive electrode

active material according to claim 7 wherein a solvent is allowed to co-exist with the precursor in said de-aerating step and is vaporized off in said inert gas to remove air contained in said precursor.

10. The method for producing a non-aqueous electrolyte secondary battery according to claim 7 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .

11. A method for producing a non-aqueous electrolyte secondary battery having a positive electrode containing a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode, said positive electrode active material being produced by a method comprising:

a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals;

a de-aerating step of de-aerating air contained in said precursor obtained in said mixing step; and

a sintering step of sintering and reacting said precursor obtained in a state free of air by said de-aerating step.

12. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein an inert gas is introduced after substituting vacuum for an atmosphere in said precursor and subsequently an inert gas is introduced to remove air contained in said precursor.

13. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein a solvent is allowed to co-exist with the precursor in said de-aerating step and is vaporized off in said inert gas to remove air contained in said precursor.

14. The method for producing a non-aqueous electrolyte secondary battery according to claim 11 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .

15. A method for producing a positive electrode active material comprising:

a mixing step of mixing a plurality of sub-

stances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals; and

a sintering step of sintering and reacting said precursor obtained by said mixing step; wherein an electrically conductive agent is added to said starting material for synthesis or said precursor.

16. The method for producing a positive electrode active material according to claim 15 wherein said electrically conductive agent is added in an amount of 0.5 to 20 parts by weight to 100 parts by weight of said  $\text{Li}_x\text{M}_y\text{PO}_4$ .

17. The method for producing a positive electrode active material according to claim 15 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ .

18. The method for producing a positive electrode active material according to claim 15 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$ .

19. The method for producing a positive electrode active material according to claim 15 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{Li}_x\text{Mn}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$  and  $y$  is such that  $0.8 \leq y \leq 1.2$ .

20. The method for producing a positive electrode active material according to claim 15 wherein said electrically conductive agent is carbon.

21. A method for producing a non-aqueous electrolyte secondary battery having a positive electrode active material capable of reversibly doping/undoping lithium, a negative electrode mounted facing said positive electrode and capable of reversibly doping/undoping lithium, and a non-aqueous electrolyte interposed between said positive electrode and the negative electrode,

wherein said positive electrode active material is synthesized by a mixing step of mixing a plurality of substances to give a precursor, said substances proving a starting material for synthesis of a compound represented by the general formula  $\text{Li}_x\text{M}_y\text{PO}_4$  where  $x$  is such that  $0 < x \leq 2$ ,  $y$  is such that  $0.8 \leq y \leq 1.2$  and  $M$  includes at least one of 3d transition metals and a sintering step of sintering and reacting said precursor obtained by said mixing step, and wherein an electrically conductive agent is added to said starting material for synthesis or to said precursor to synthesize said positive electrode active material.

22. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said electrically conductive agent is added in an amount of 0.5 to 20 parts by weight to 100 parts by weight of said  $\text{Li}_x\text{M}_y\text{PO}_4$ . 5
23. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFePO}_4$ . 10
24. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{LiFe}_x\text{Mn}_{1-x}\text{PO}_4$ .
25. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said  $\text{Li}_x\text{M}_y\text{PO}_4$  is  $\text{Li}_x\text{Mn}_y\text{PO}_4$  where x is such that  $0 < x \leq 2$  and y is such that  $0.8 \leq y \leq 1.2$ . 15
26. The method for producing a non-aqueous electrolyte secondary battery according to claim 21 wherein said electrically conductive agent is carbon. 20

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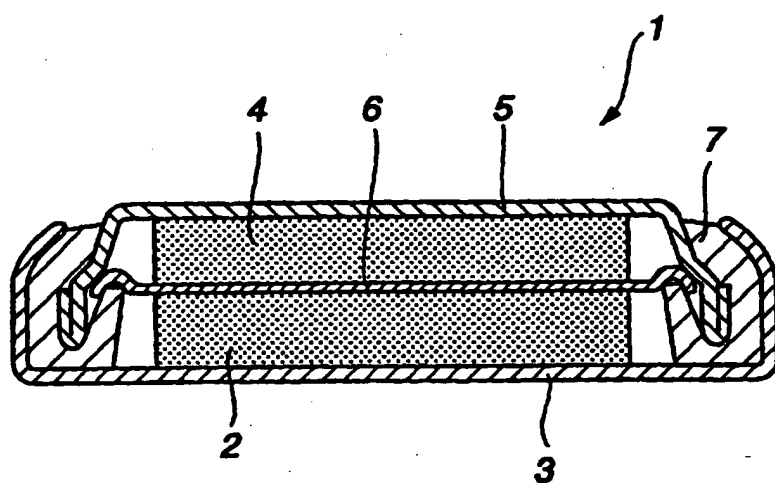
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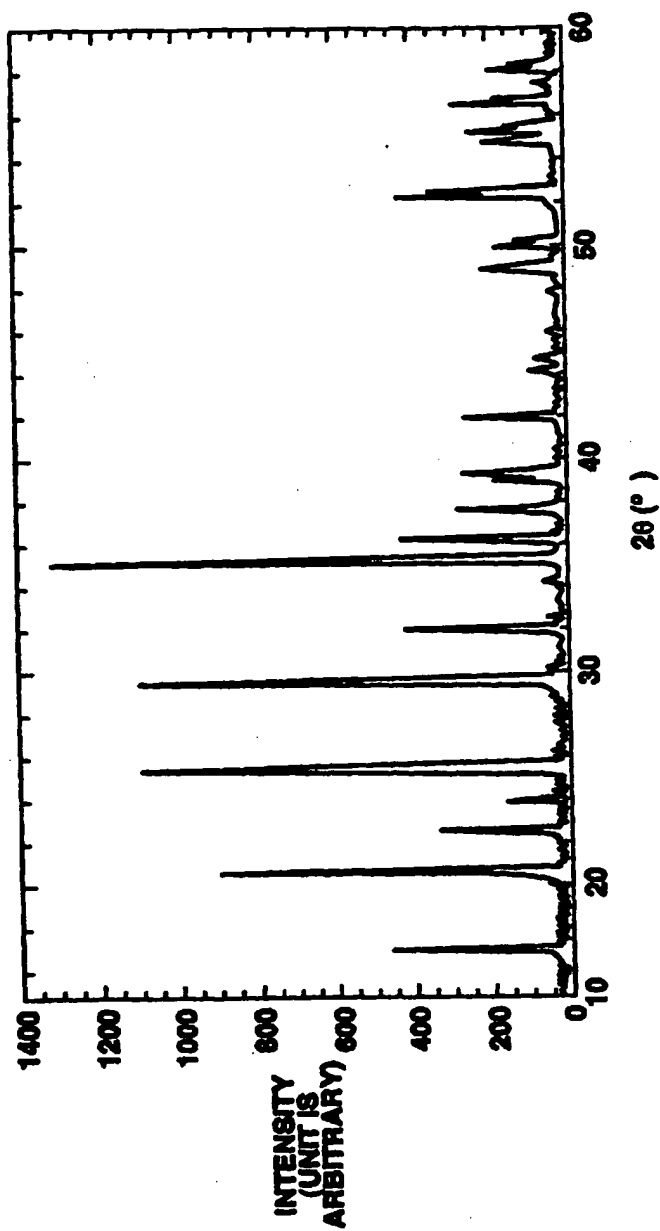
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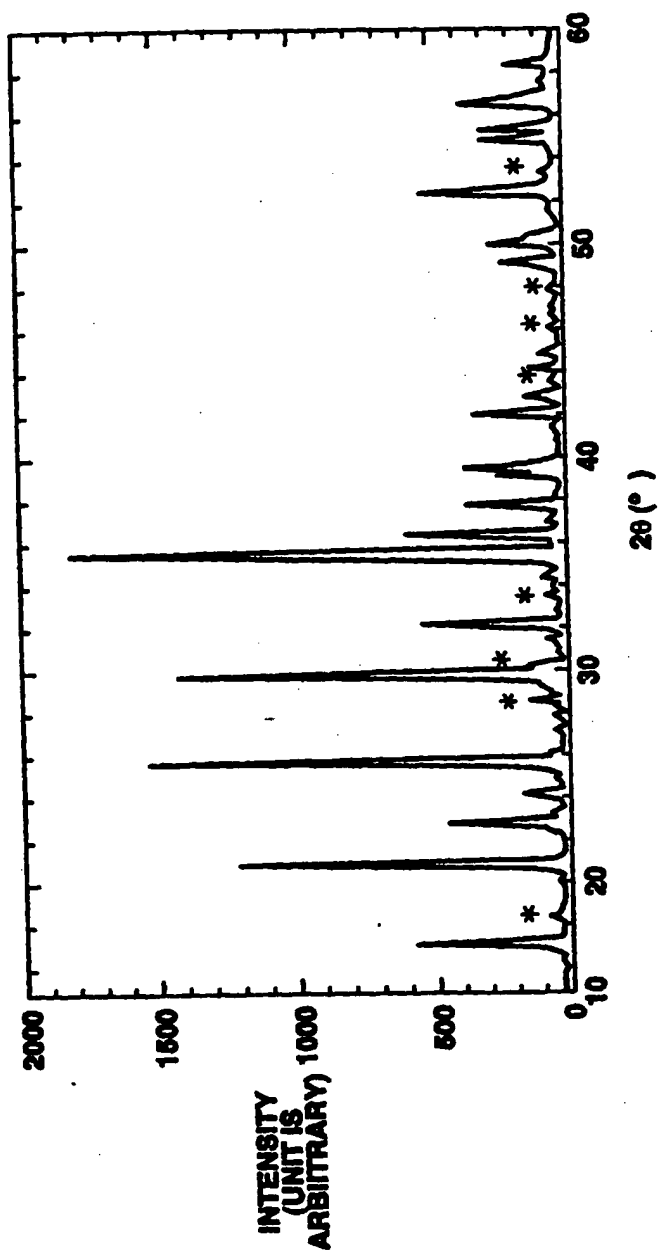
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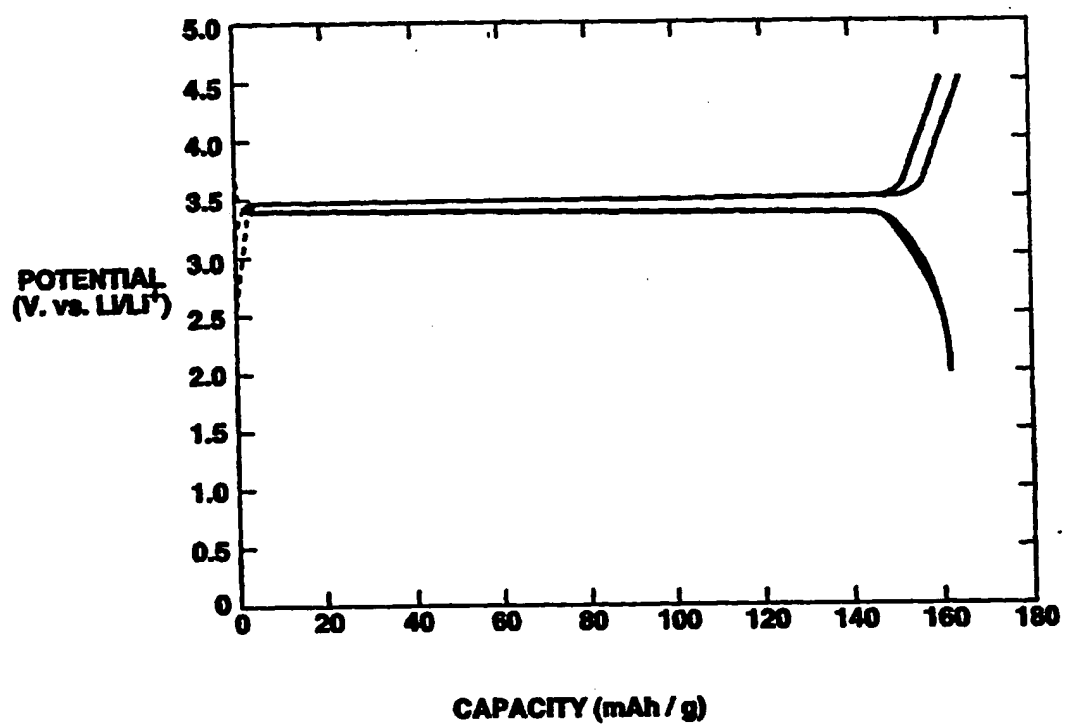
**FIG.1**



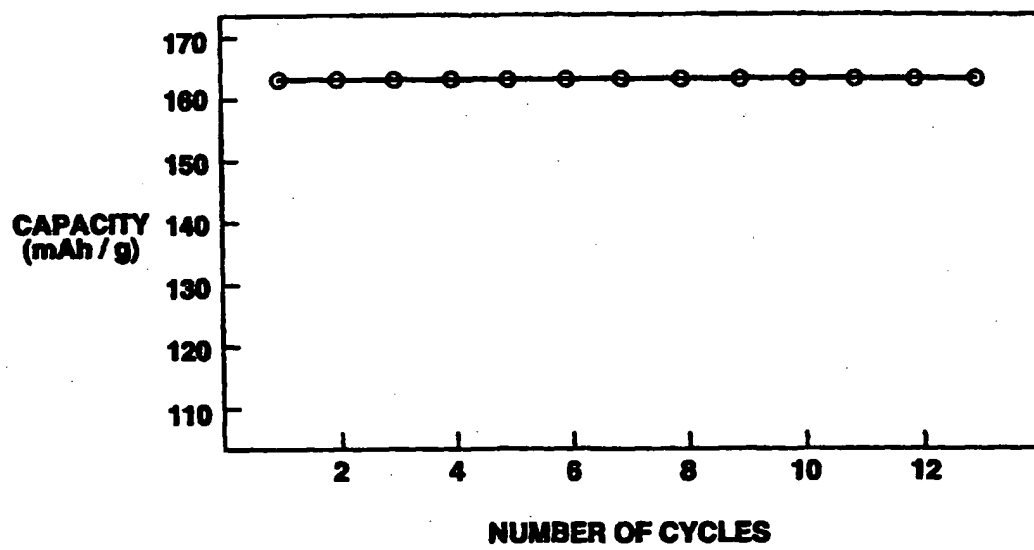
**FIG.2**



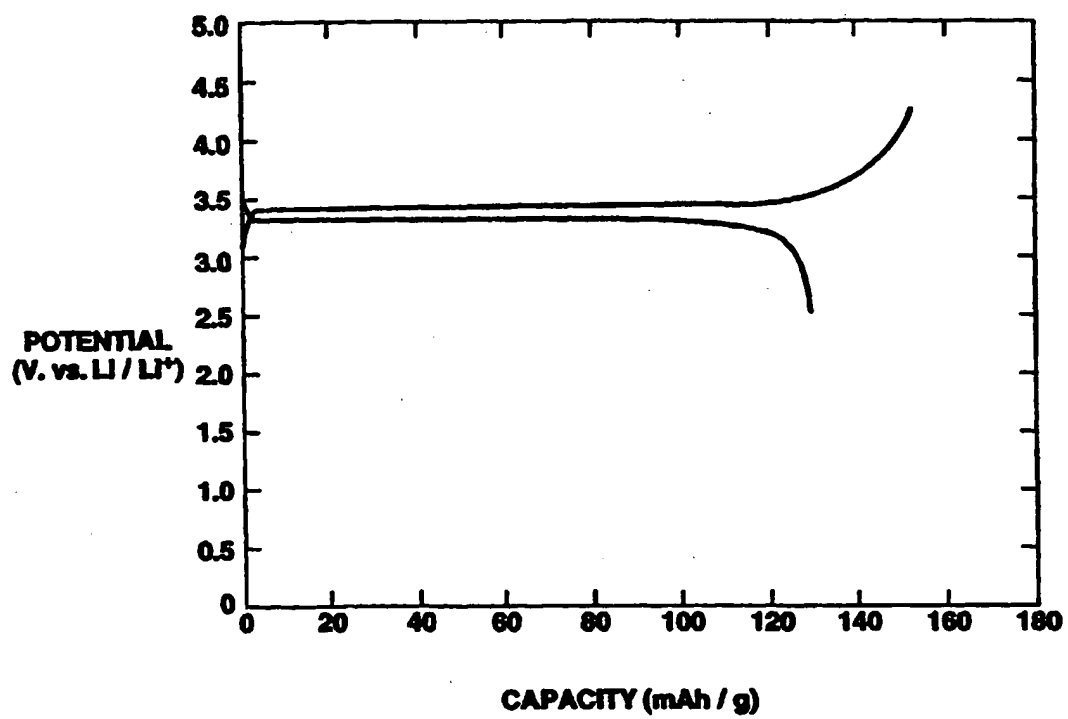
**FIG.3**

**FIG.4**

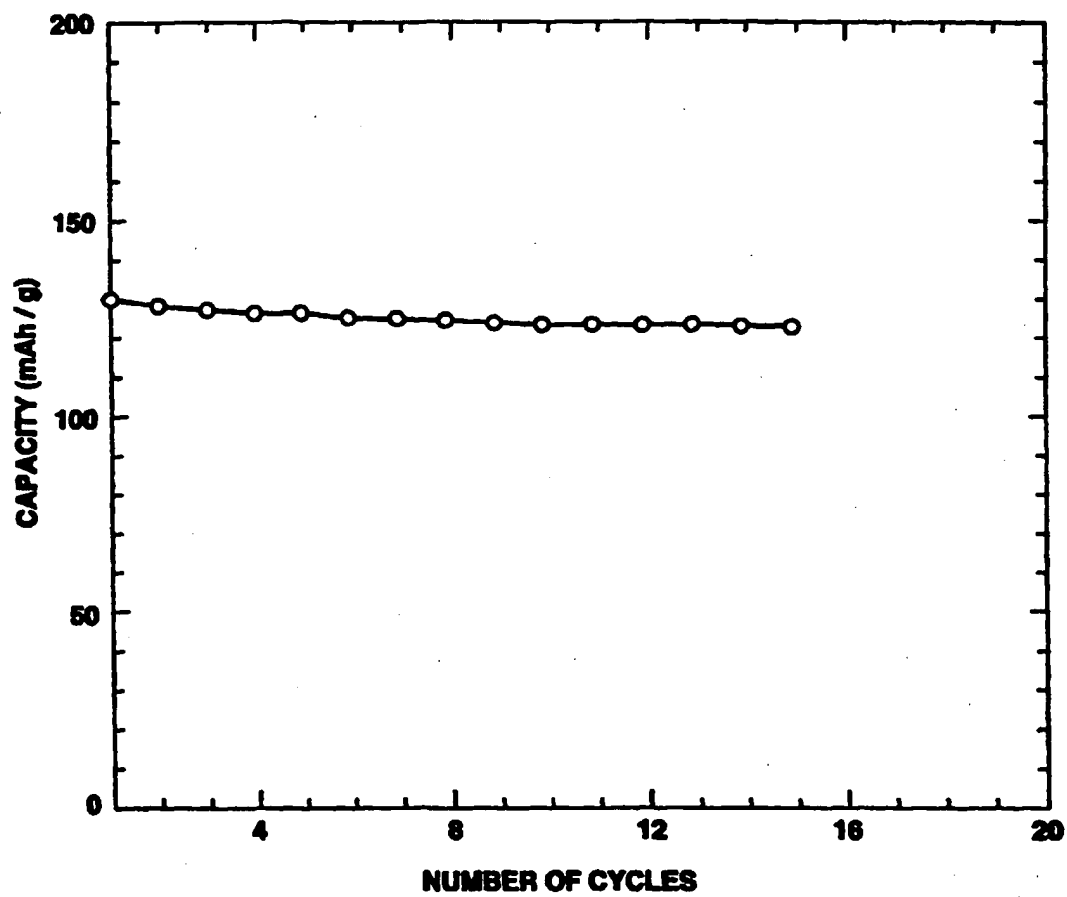




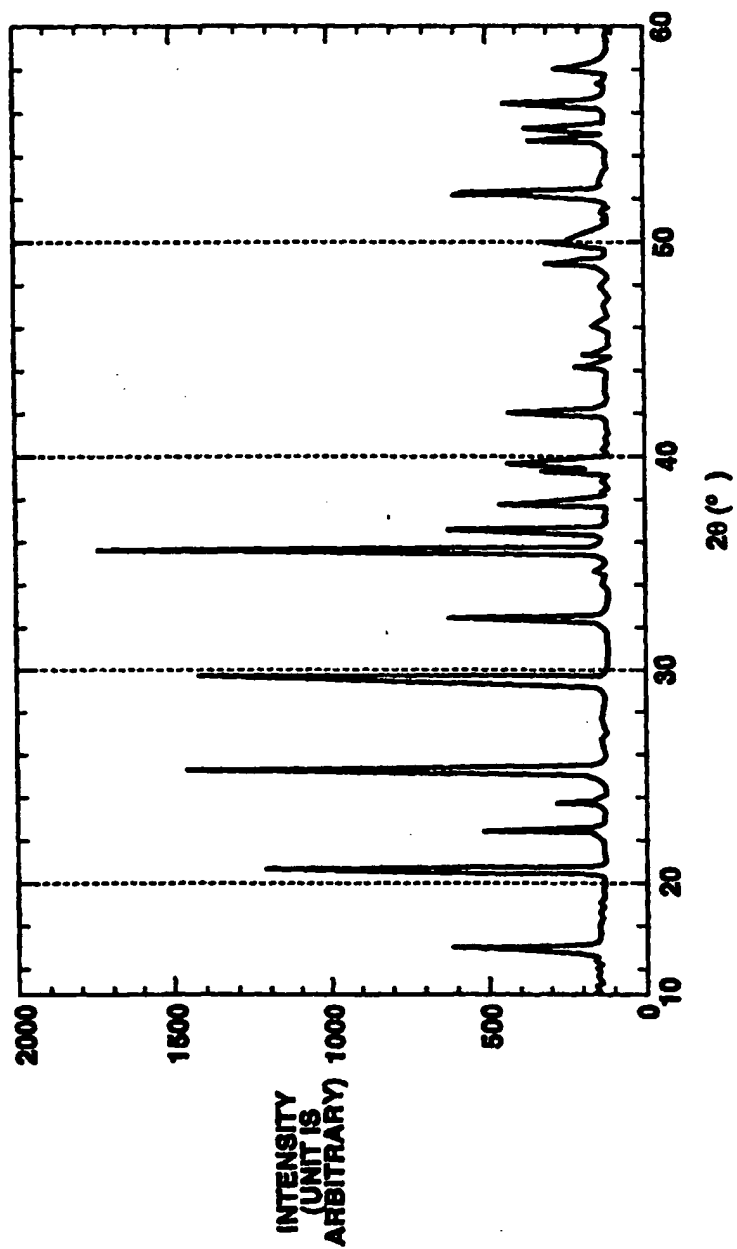
**FIG.5**



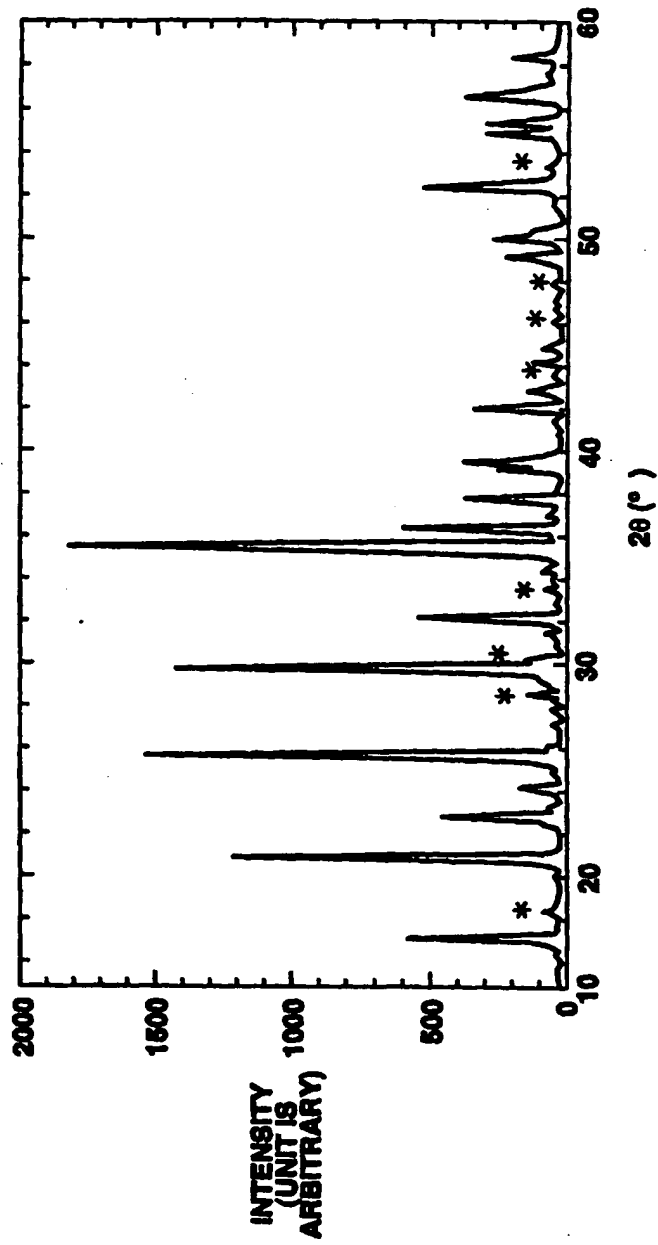
**FIG.6**



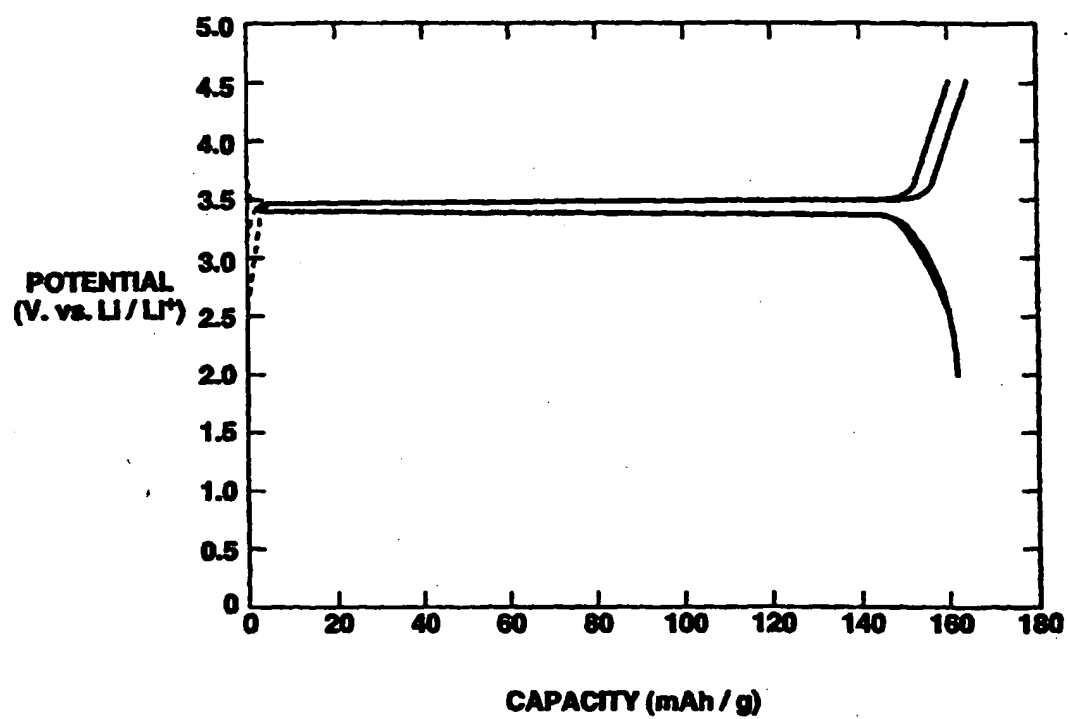
**FIG.7**

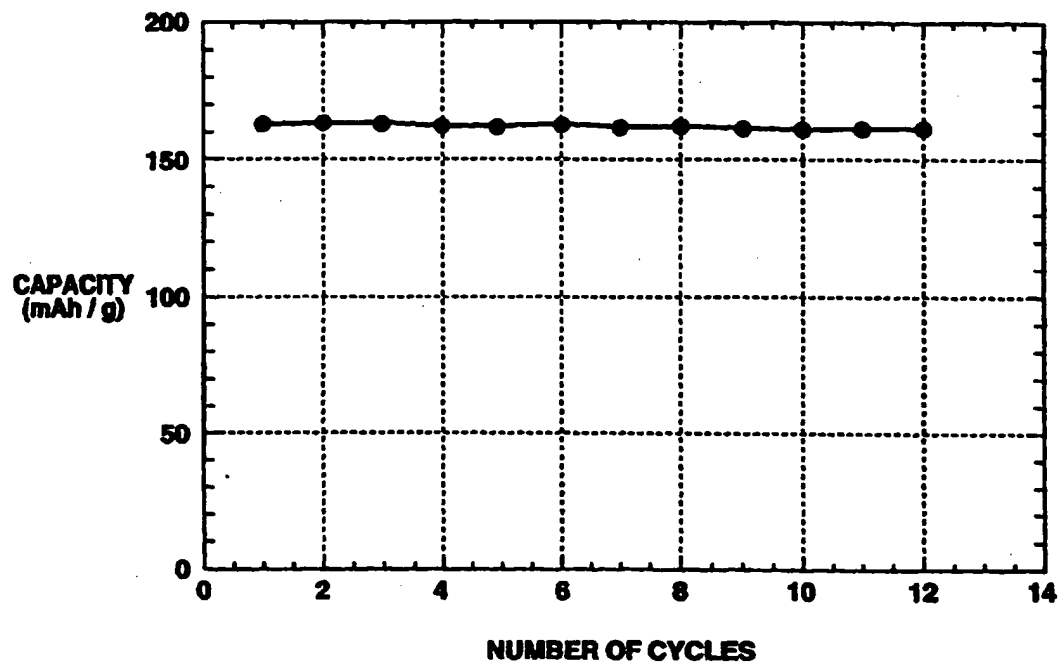


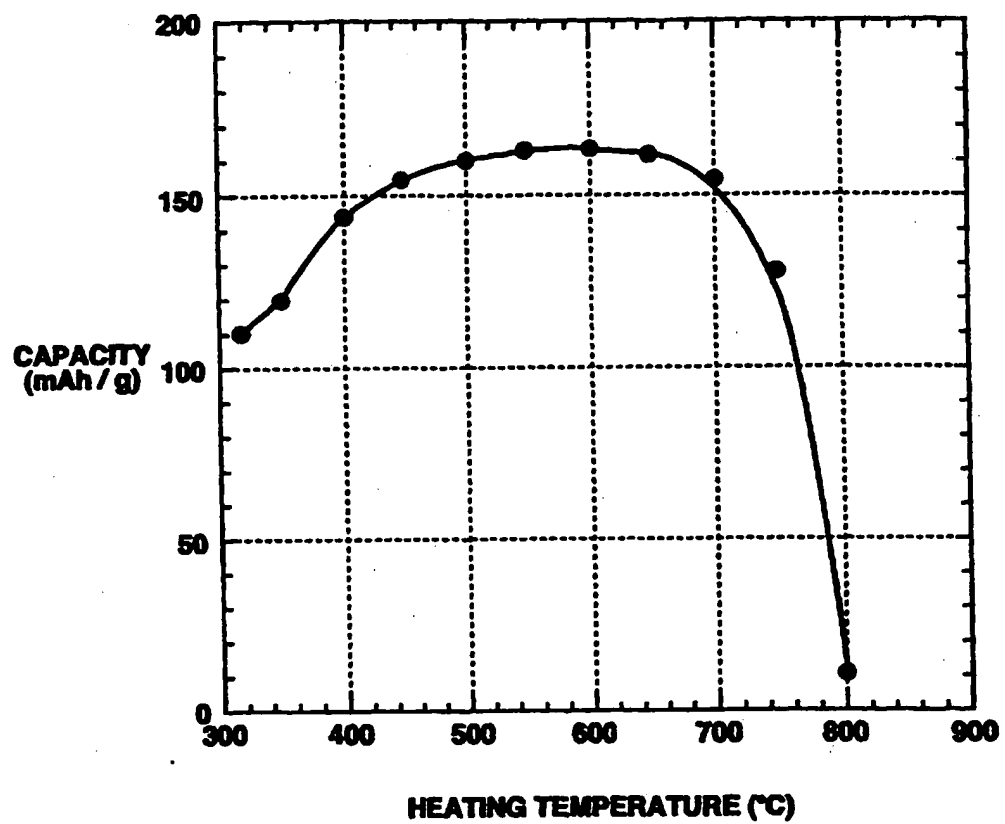
**FIG.8**



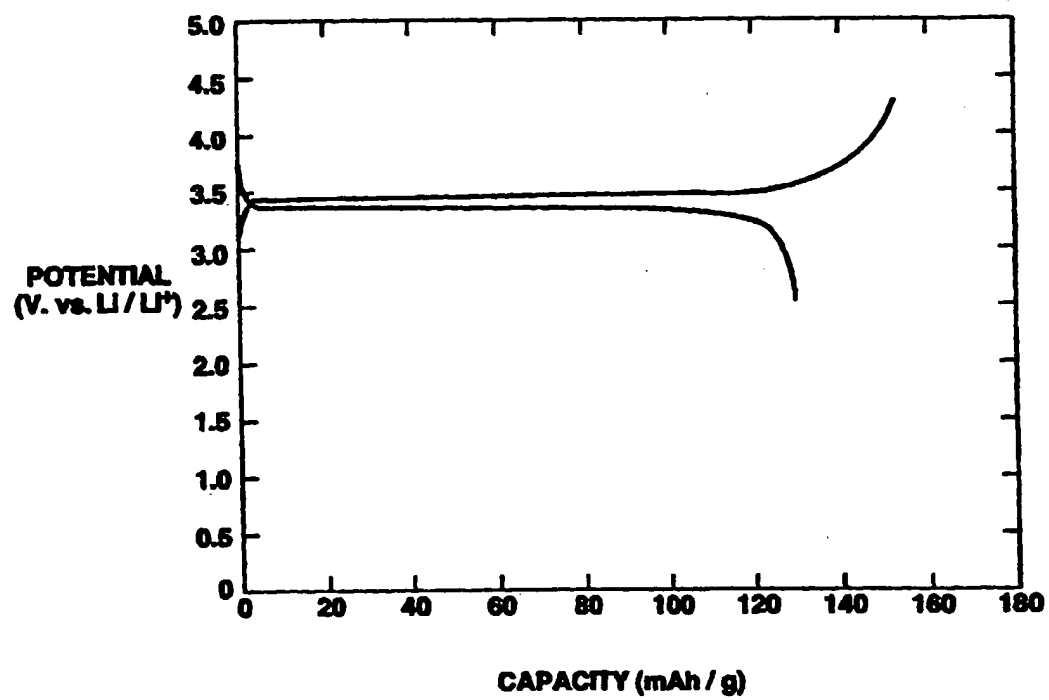
**FIG.9**

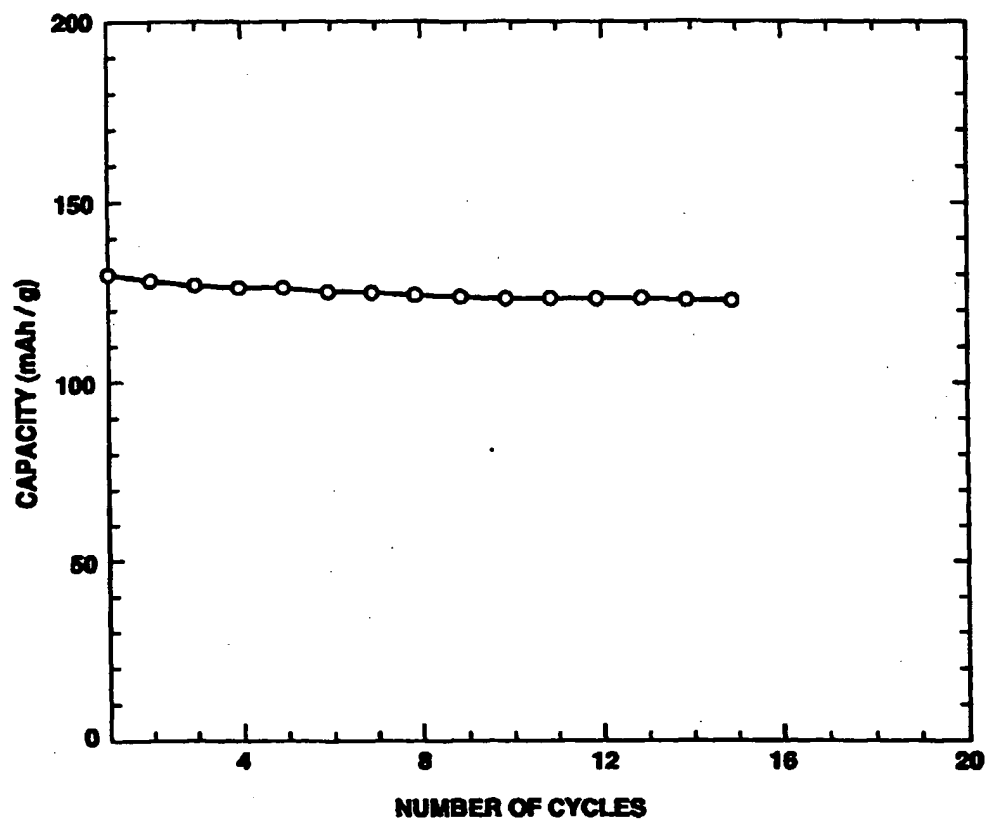
**FIG.10**

**FIG.11**

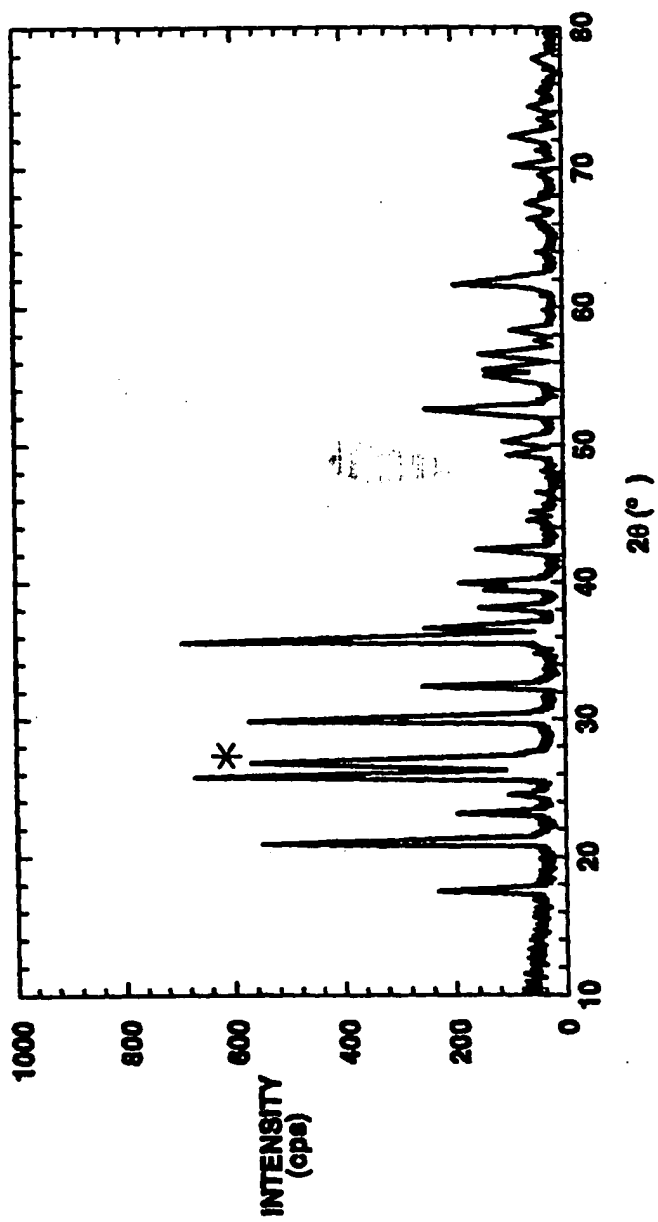
**FIG.12**



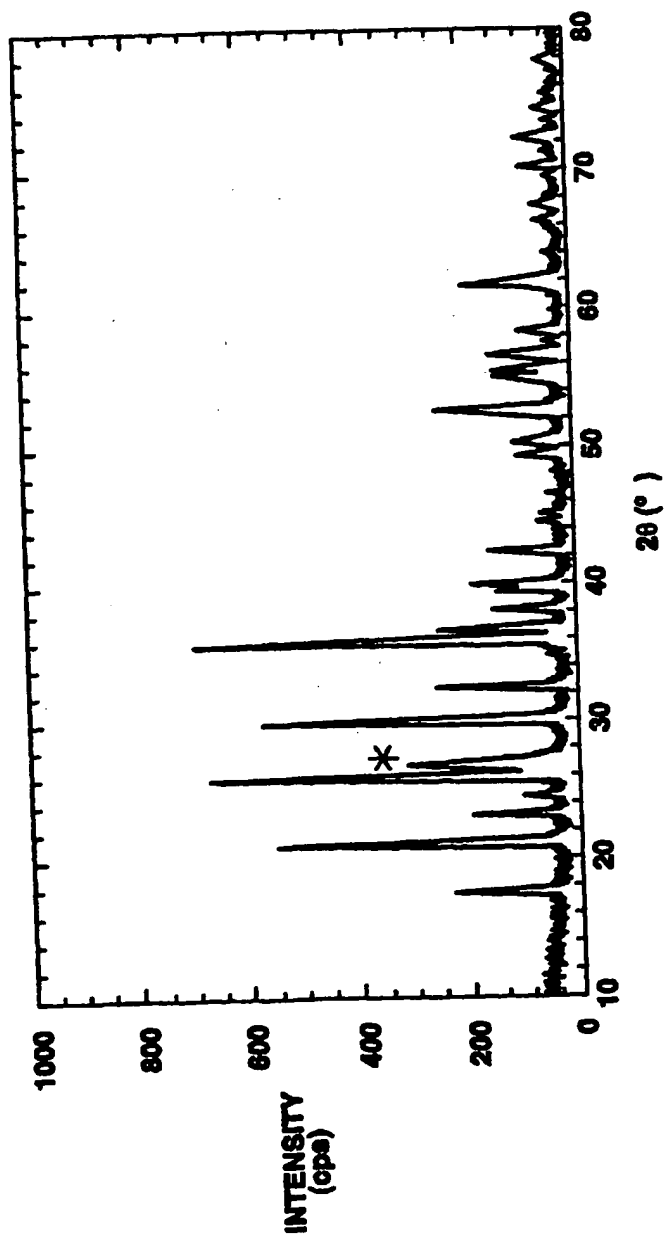
**FIG.13**



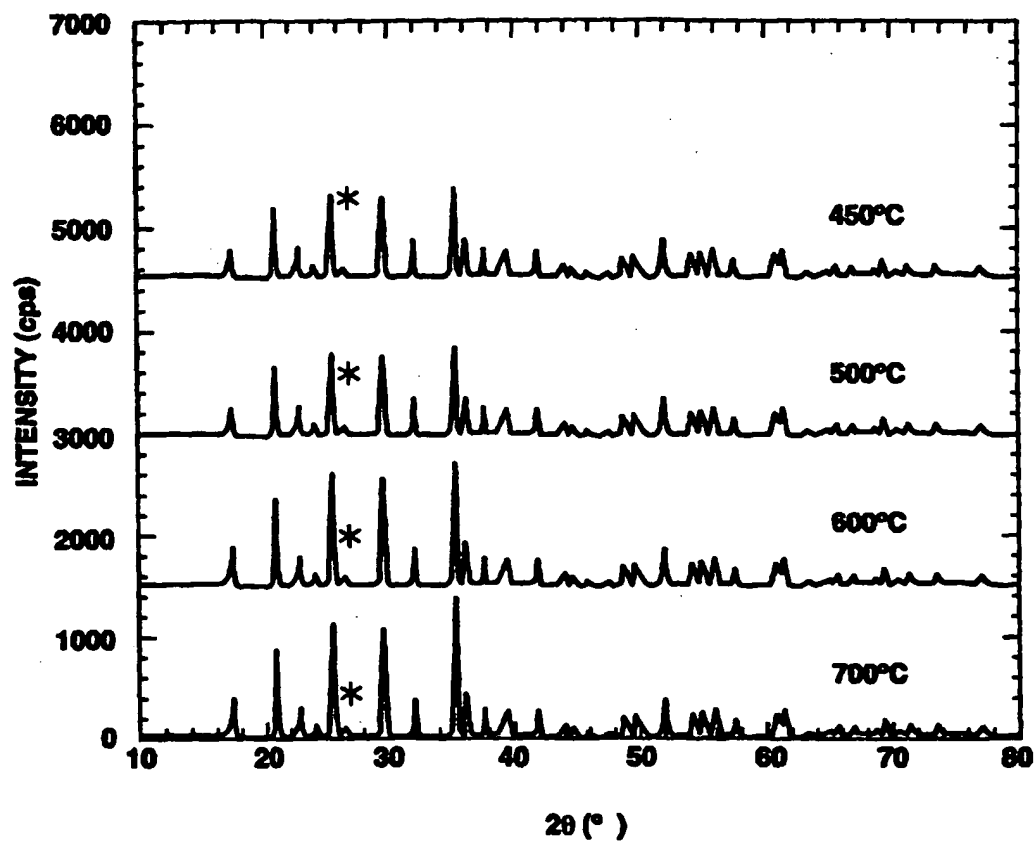
**FIG.14**

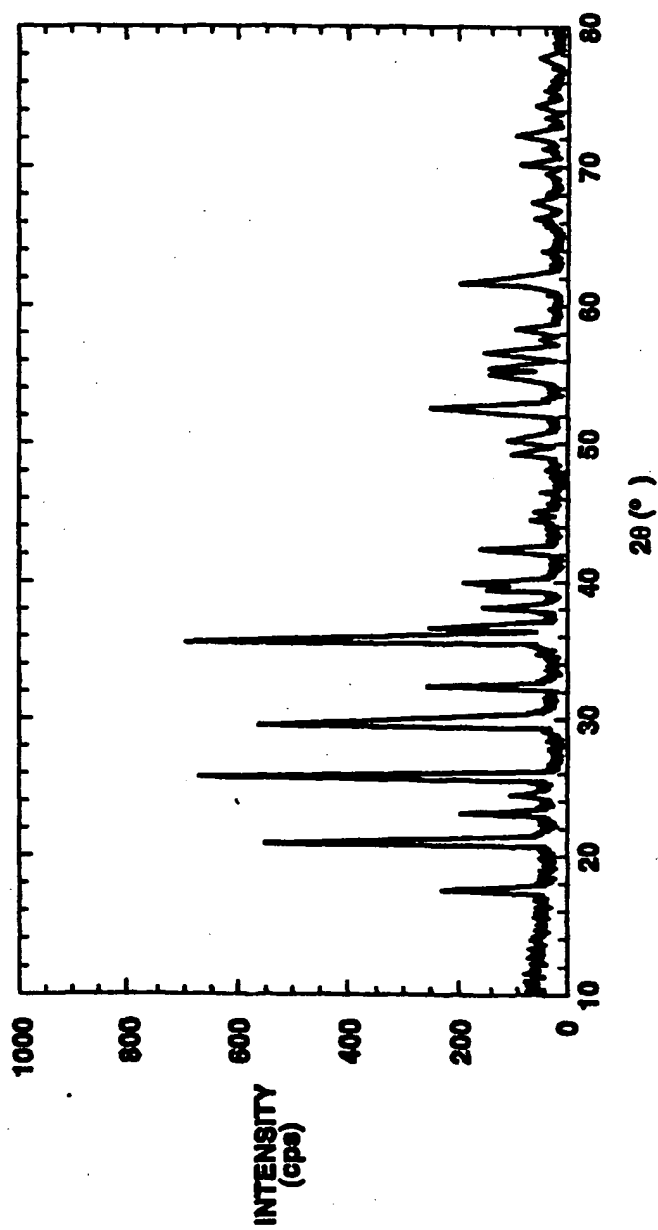


**FIG.15**

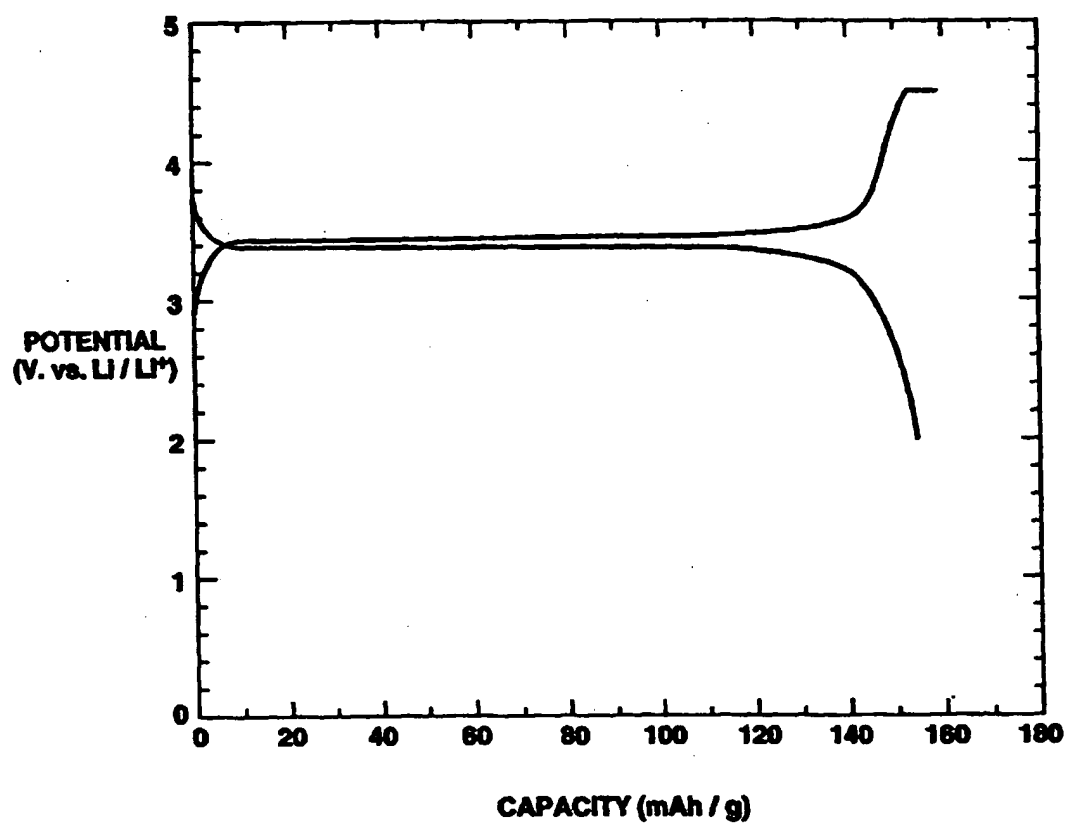


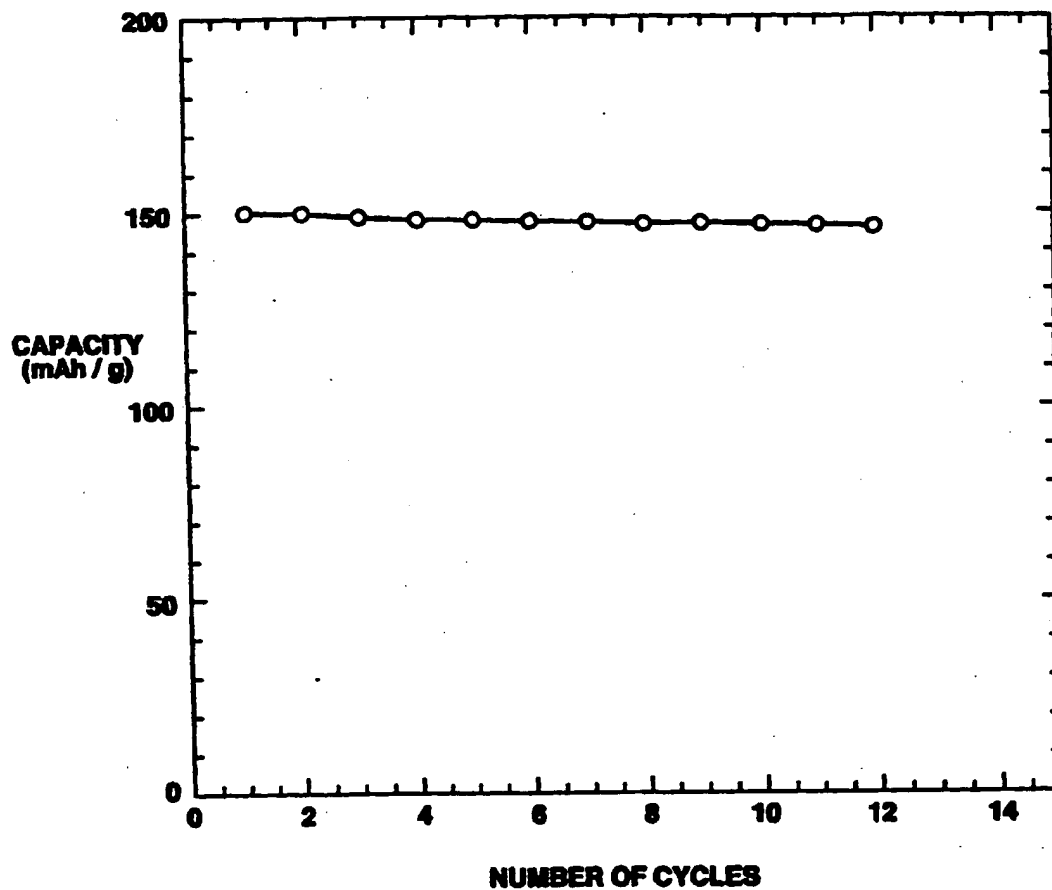
**FIG.16**

**FIG.17**



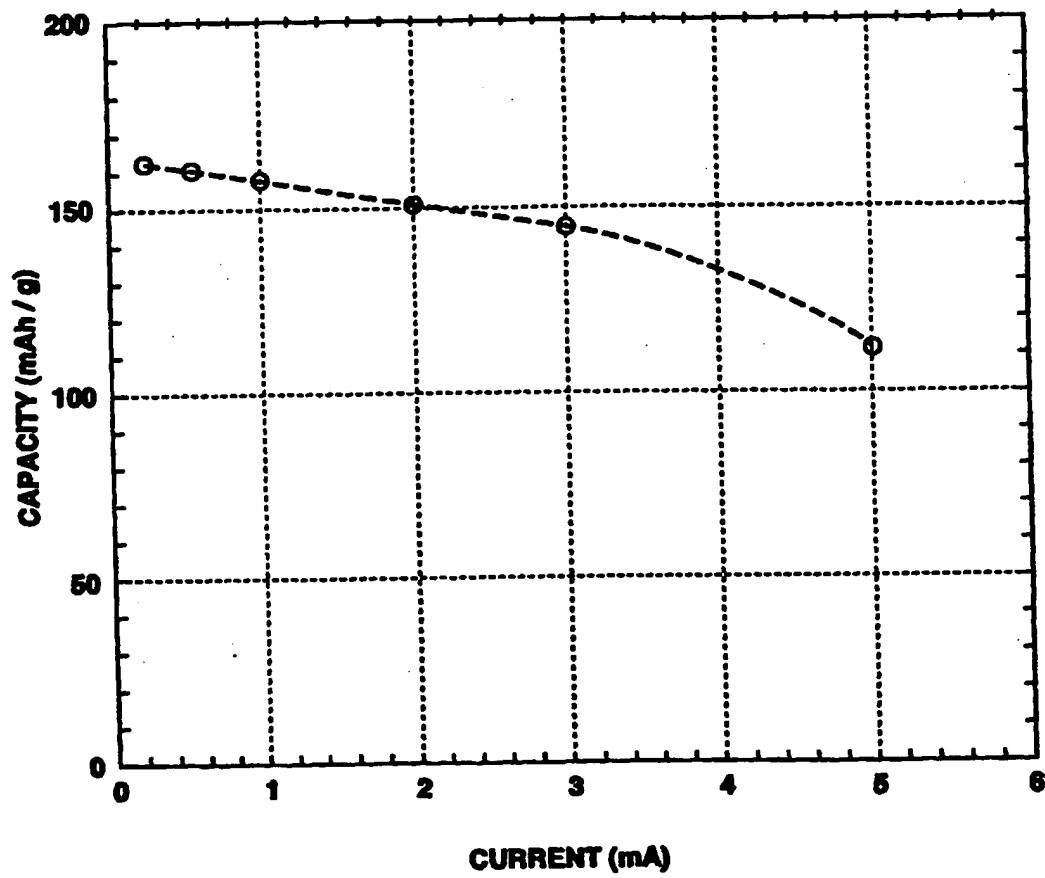
**FIG.18**

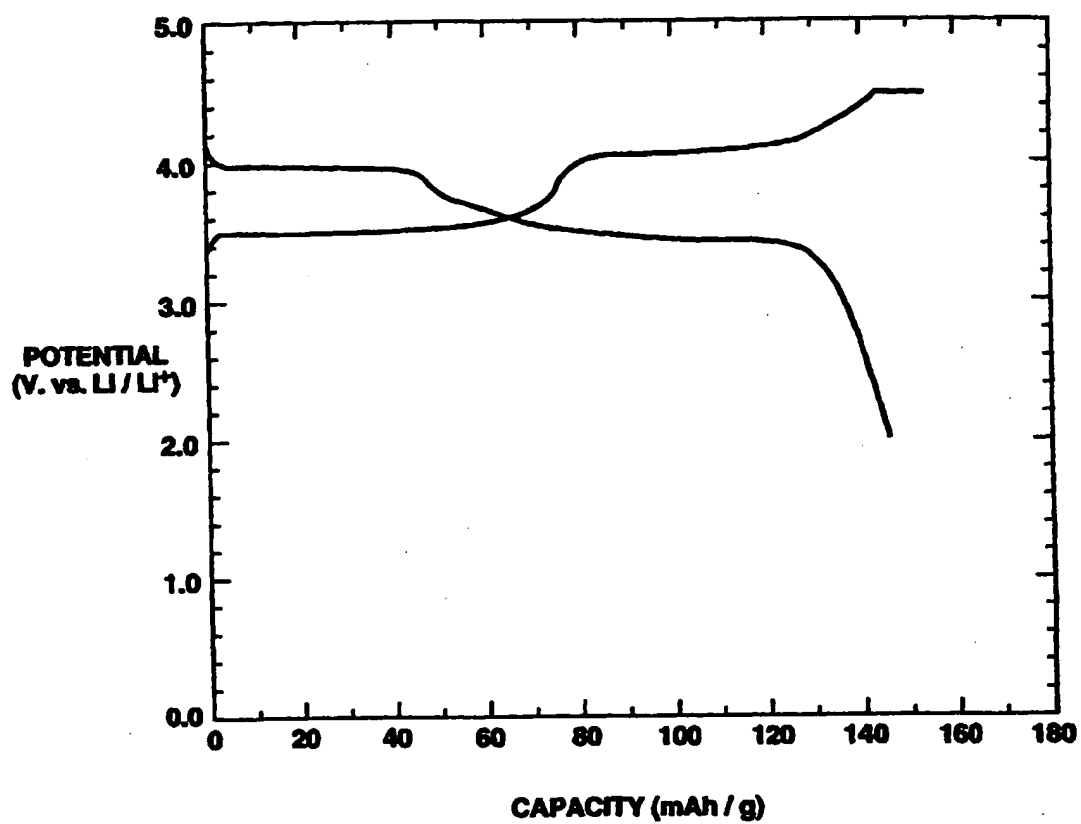
**FIG.19**



**FIG.20**



**FIG.21**

**FIG.22**

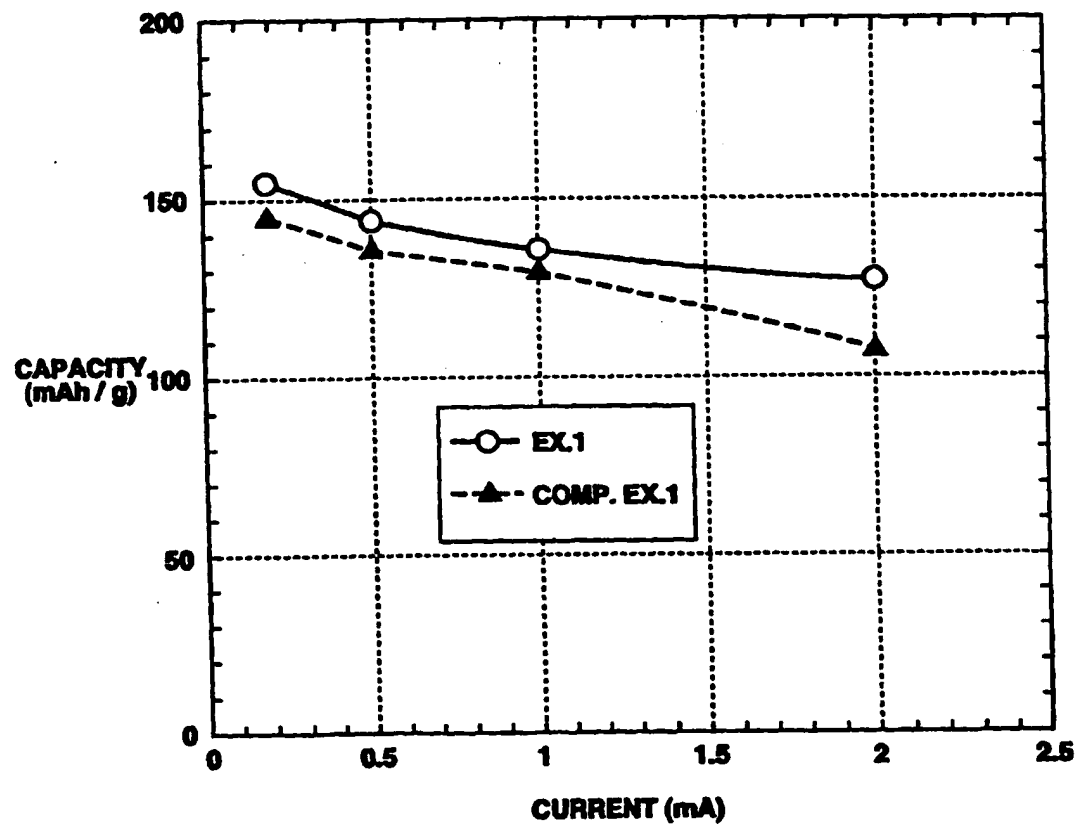
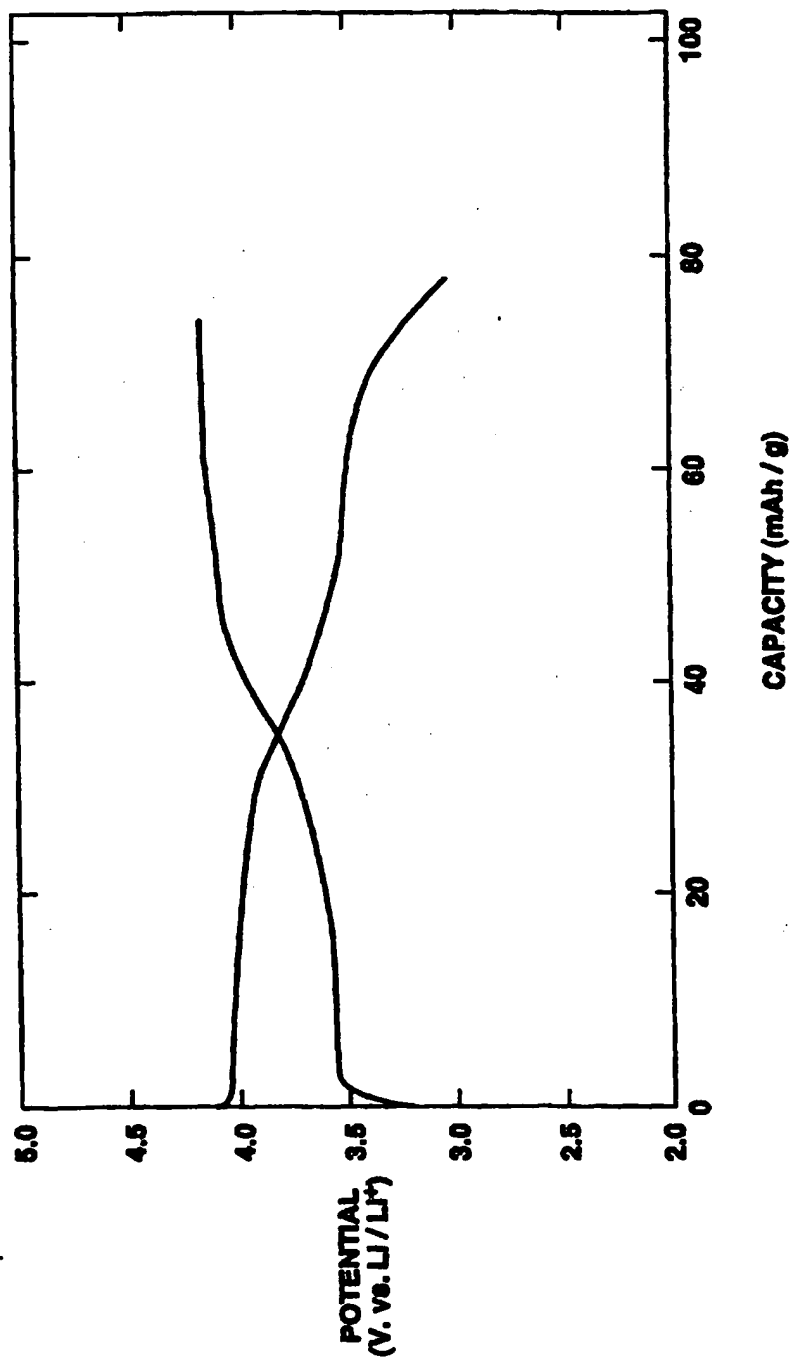
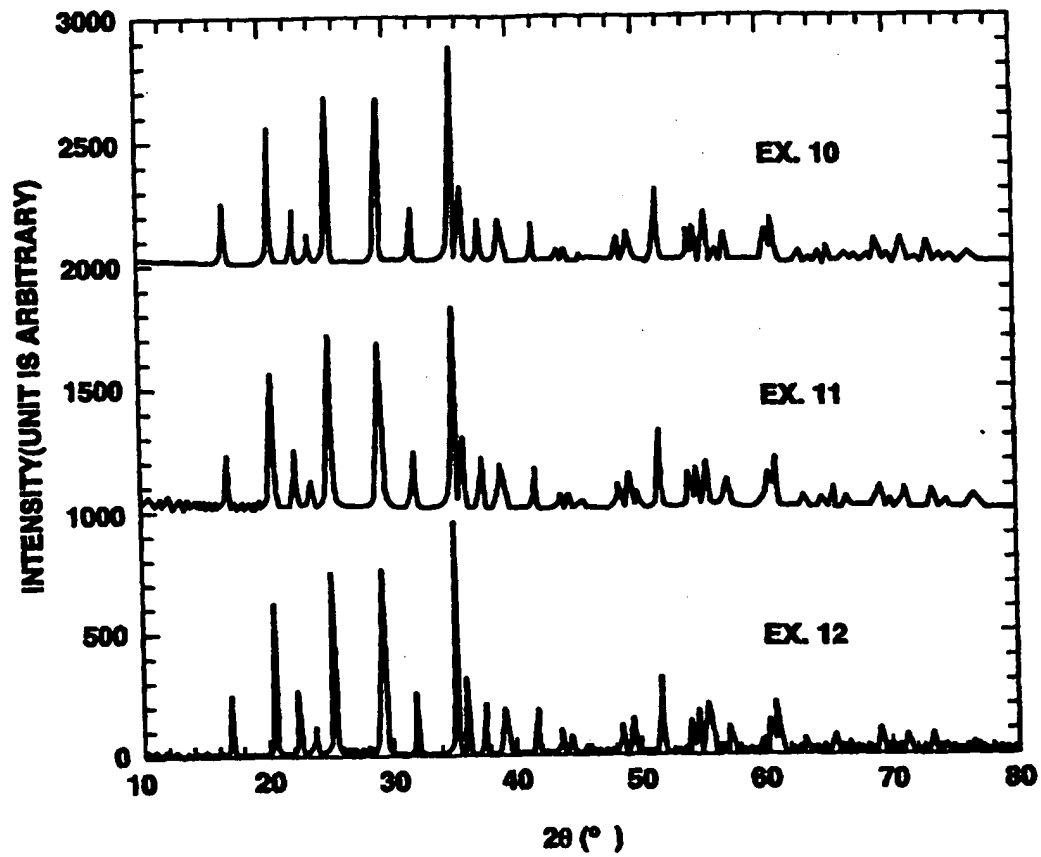


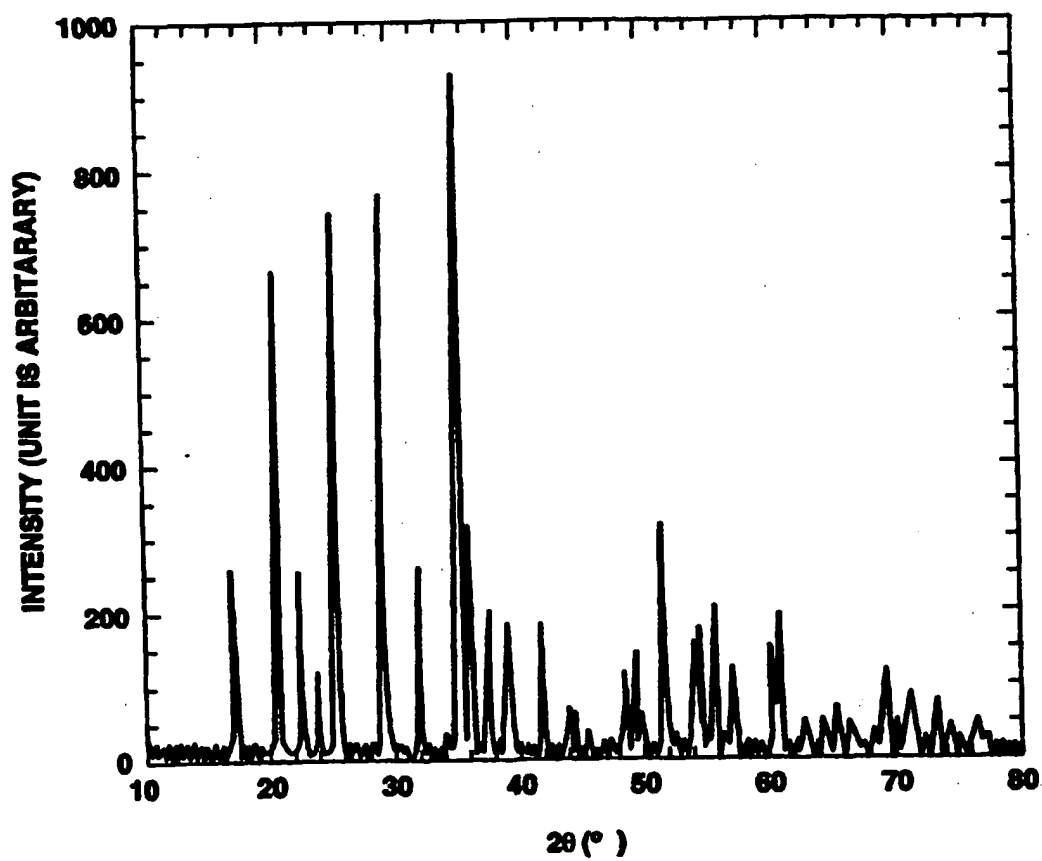
FIG.23



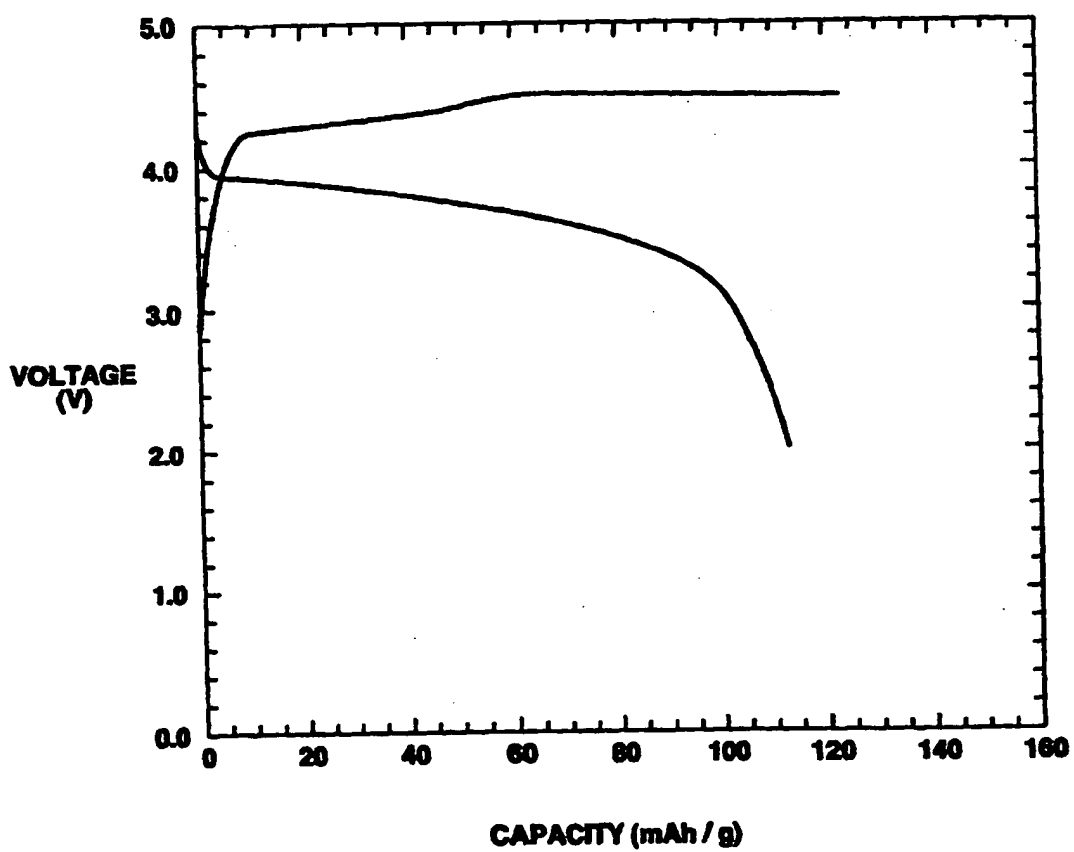
**FIG.24**



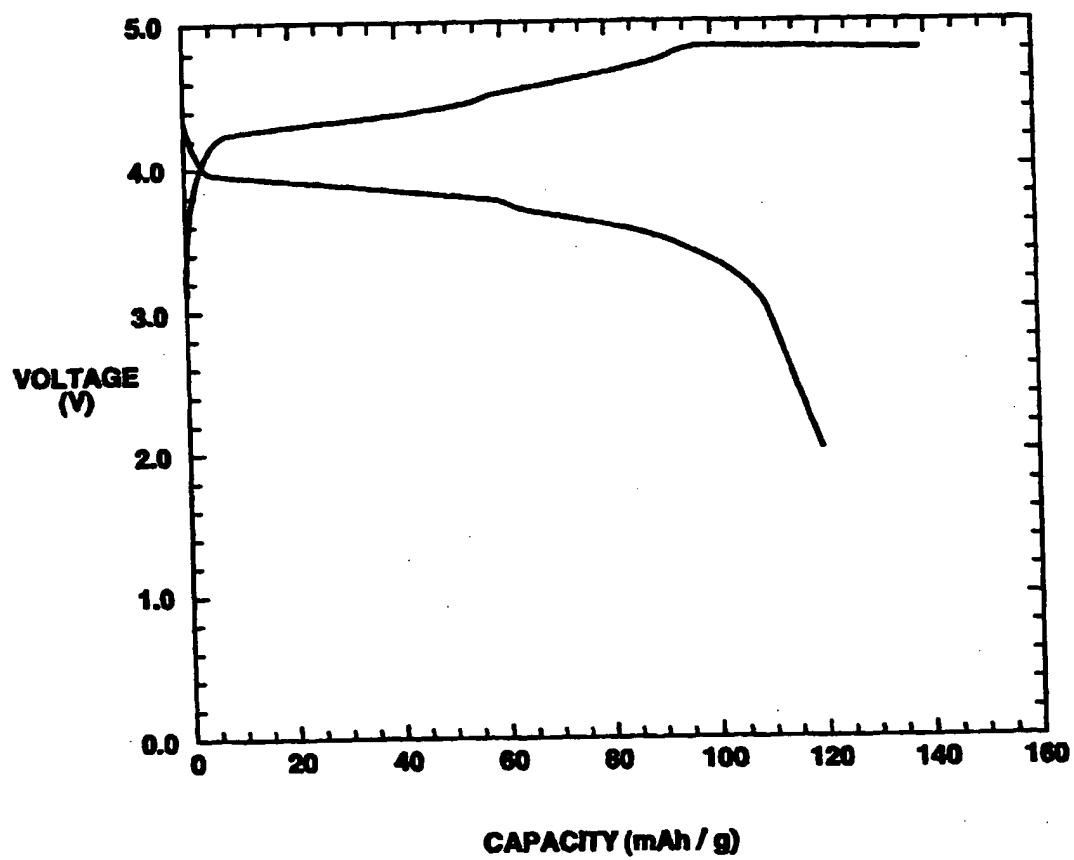
**FIG.25**



**FIG.26**

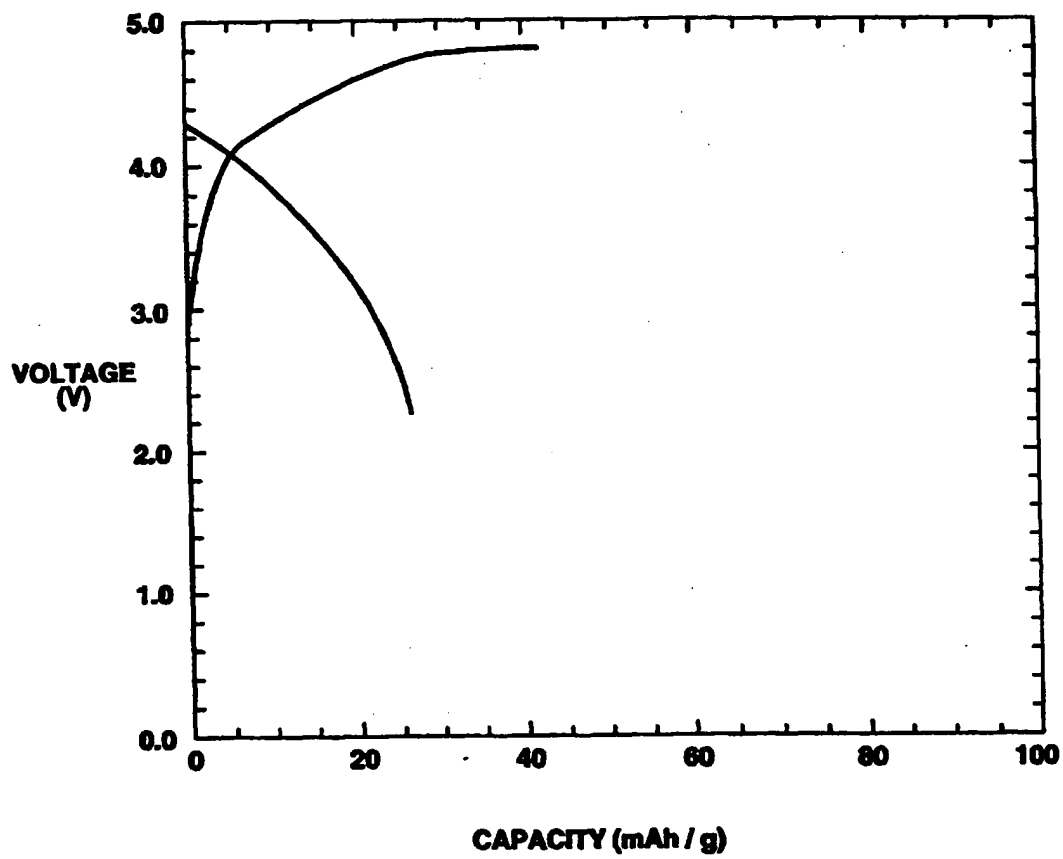


**FIG.27**

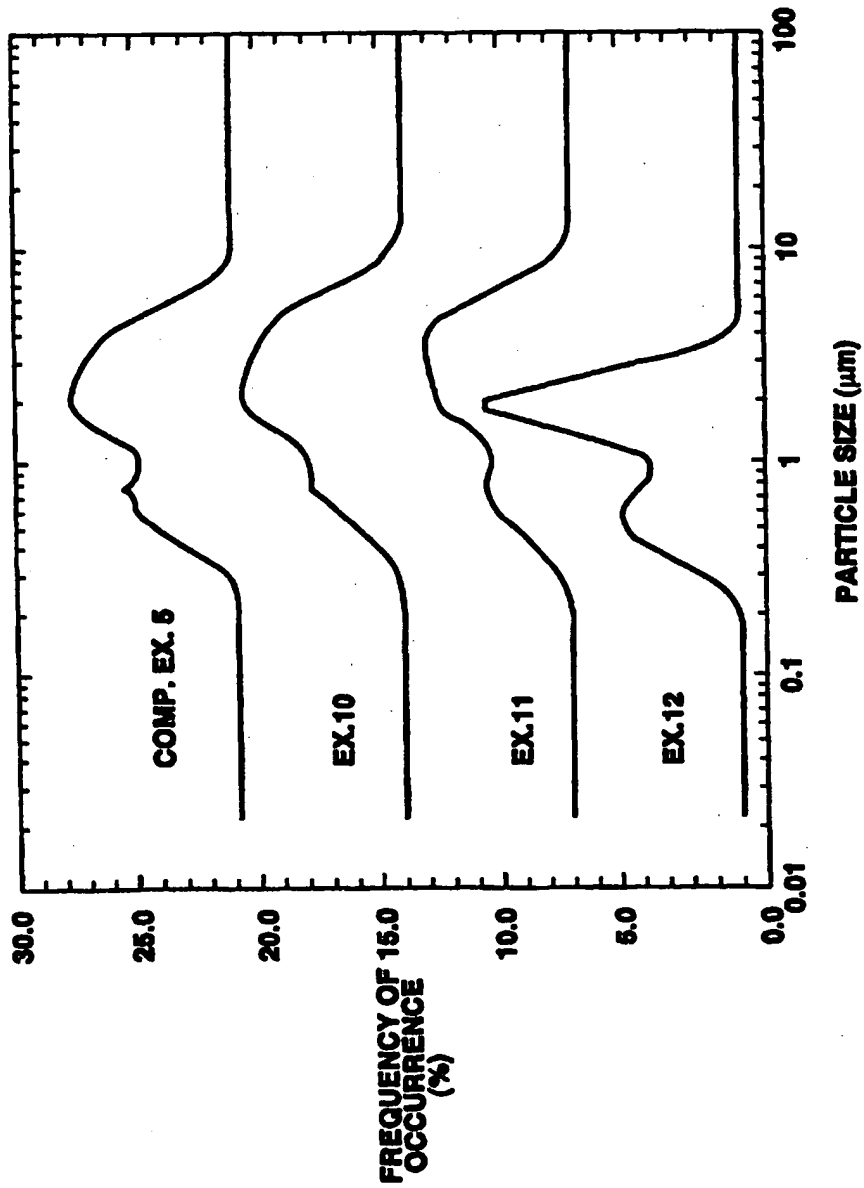


**FIG.28**

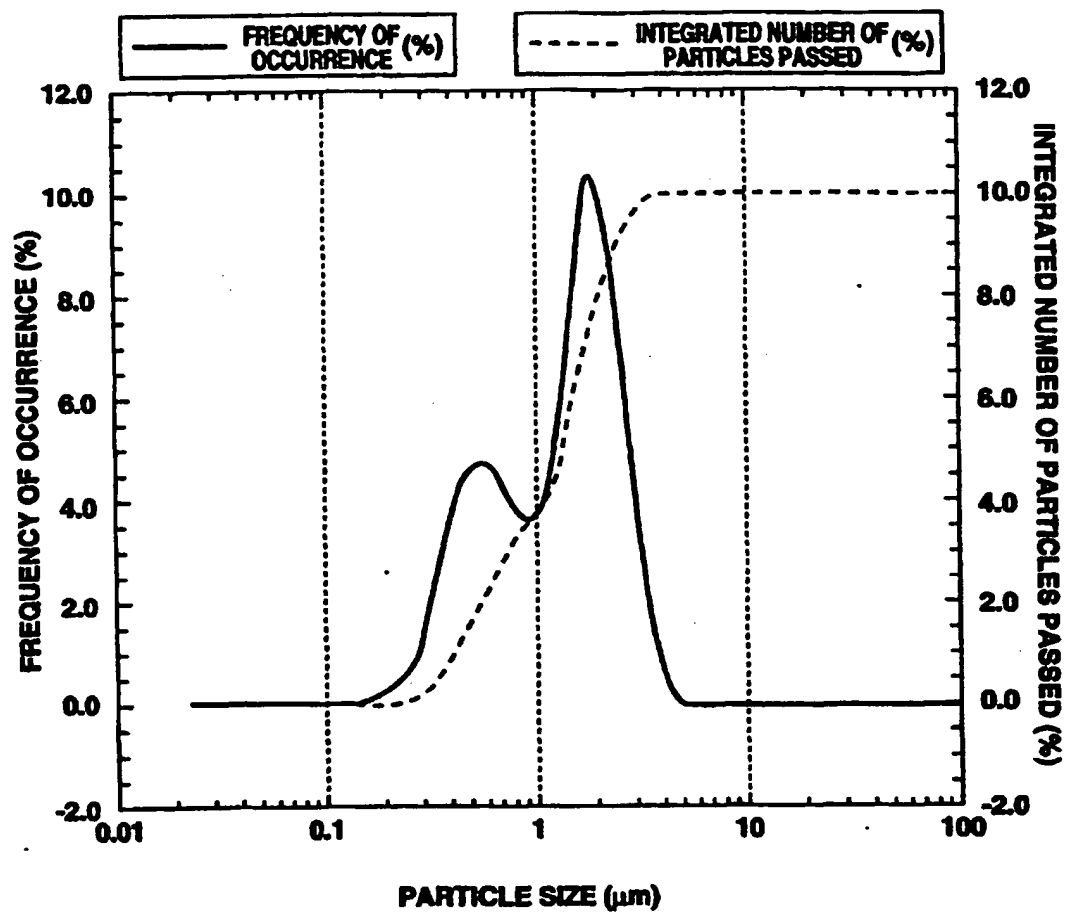




**FIG.29**



**FIG.30**

**FIG.31**

## INTERNATIONAL SEARCH REPORT

International application No.

PCT/JP00/01915

<b>A. CLASSIFICATION OF SUBJECT MATTER</b> Int.Cl. <sup>7</sup> H01M4/58, H01M10/40, H01M4/04 According to International Patent Classification (IPC) or to both national classification and IPC		
<b>B. FIELDS SEARCHED</b> Minimum documentation searched (classification system followed by classification symbols) Int.Cl. <sup>7</sup> H01M4/02, 4/04, 4/36-4/62, 10/36-10/40 Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched Jitsuyo Shinan Koho 1926-1996 Jitsuyo Shinan Toroku Koho 1996-2000 Kokai Jitsuyo Shinan Koho 1971-2000 Toroku Jitsuyo Shinan Koho 1994-2000 Electronic data base consulted during the international search (name of data base and, where practicable, search terms used) JICST		
<b>C. DOCUMENTS CONSIDERED TO BE RELEVANT</b>		
Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	JP, 10-312789, A (YUASA CORPORATION), 24 November, 1998 (24.11.98) (Family: none)	1-26
A	JP, 9-171827, A (Sony Corporation), 30 June, 1997 (30.06.97) (Family: none)	1-26
A	JP, 9-134724, A (Nippon Telegr. & Teleph. Corp. <NTT>), 20 May, 1997 (20.05.97) (Family: none)	1-26
A	JP, 6-283207, A (Nippon Telegr. & Teleph. Corp. <NTT>), 07 October, 1994 (07.10.94) (Family: none)	1-26
<input type="checkbox"/> Further documents are listed in the continuation of Box C. <input type="checkbox"/> See patent family annex.		
* Special categories of cited documents: "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed		"I" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family
Date of the actual completion of the international search 06 June, 2000 (06.06.00)		Date of mailing of the international search report 20 June, 2000 (20.06.00)
Name and mailing address of the ISA/ Japanese Patent Office Facsimile No.		Authorized officer Telephone No.

Form PCT/ISA/210 (second sheet) (July 1992)